

Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point. Express the heat in kilojoules to three significant figures.

$$q = 5830 \text{ kJ}$$

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Correct

At its boiling point, the enthalpy of vaporization (ΔH_{vap}) of water is 40.7 kJ/mol. Once the given mass, $2.58 \times 10^3 \text{ g}$, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by ΔH_{vap} yields the amount of heat required to evaporate that amount of water.

Part B

Suppose that 0.48 g of water at 25°C condenses on the surface of a 55-g block of aluminum that is initially at 25°C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum, $C_{s,\text{Al}}$, is 0.903 J/(g · °C).)

Express the temperature in degrees Celsius to two significant figures.

$$T_f =$$

°C

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Mastering Chemistry Answers Chapter 12

John S. Phillips, Cheryl Wistrom



Mastering Chemistry Answers Chapter 12:

CBSE Class 12 Mastering Organic Reactions MCQ With Answers Priti Singhal, 2024-11-11 The primary objective of this book is to serve as a comprehensive guide for students educators and researchers by focusing on reaction mechanisms practical applications and problem solving techniques Organic chemistry is not just about memorizing equations and formulas it is about understanding how molecules interact change and influence each other under different conditions With that in mind this book emphasizes the logic and patterns behind organic reactions making it easier for readers to apply concepts across a variety of scenarios Each chapter of this book builds upon foundational knowledge ensuring a progressive learning experience From nucleophilic substitutions to pericyclic reactions and from oxidation reduction mechanisms to named reactions we cover both fundamental and advanced topics to cater to students at all levels Real world examples have been integrated throughout the chapters to show how organic reactions play essential roles in pharmaceuticals biochemistry agriculture and environmental science This approach bridges the gap between theory and practical applications helping readers appreciate the relevance of organic chemistry in daily life

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