

Short Answer Type Questions

1. What will be the mass of one atom of C-12 in grams?
2. How many significant figures should be present in the answer of the following calculations?

$$\frac{2.5 \times 1.25 \times 3.5}{2.01}$$

3. What is the symbol for SI unit of mole? How is the mole defined?
4. What is the difference between molality and molarity?
5. Calculate the mass percent of calcium, phosphorus and oxygen in calcium phosphate $\text{Ca}_3(\text{PO}_4)_2$.
6. 45.4 L of dinitrogen reacted with 22.7 L of dioxygen and 45.4 L of nitrous oxide was formed. The reaction is given below:
$$2\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{N}_2\text{O}(\text{g})$$

Which law is being obeyed in this experiment? Write the statement of the law?
7. If two elements can combine to form more than one compound, the masses of one element that combine with a fixed mass of the other element, are in whole number ratio.
 - (a) Is this statement true?
 - (b) If yes, according to which law?
 - (c) Give one example related to this law.

8. Calculate the average atomic mass of hydrogen using the following data :

Isotope	% Natural abundance	Molar mass
^1H	99.985	1
^2H	0.015	2

9. Hydrogen gas is prepared in the laboratory by reacting dilute HCl with granulated zinc.

Following reaction takes place.



Calculate the volume of hydrogen gas liberated at STP when 32.65 g of zinc reacts with HCl. 1 mol of a gas occupies 22.7 L volume at STP; atomic mass of Zn = 65.3 u.

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SJ Ball



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