

Homework 1

The average human lung expands by about 0.50 LL during each breath.	Work = $-p \times \Delta$ $= -(1.0) \times (0.50\text{L})\text{atm}$ $= -0.50\text{L}\times\text{atm}$ Conversion of $1\text{L}\times\text{atm} \rightarrow \text{J} = /101.325$ $= -50.66$ or rounded to -51
The air within a piston equipped with a cylinder absorbs 545 J of heat and expands from an initial volume of 0.11 LL to a final volume of 0.88 LL against an external pressure of 1.0 barbar.	$W = -p \times \Delta V$ $\Delta V = 0.88 - 0.11 = .77\text{L}$ $= -1 \times 0.77$ $= -77\text{L} \times \text{atm} \times 101.325 = -78.02\text{J}$ $\Delta E = \text{change in internal energy}$ $q = \text{heat} = 545\text{J}$ $w = 78.02$ $\Delta E = q + w$ $= 545 + (-78.02)$ $= 466.98\text{J}$ $= 467\text{J}$
A gas is compressed from an initial volume of 5.35 LL to a final volume of 1.23 LL by an external pressure of 1.00 barbar . During the compression, the gas releases 121 J of heat.	$\Delta E = q + w$ $q = -121$ Calc work done $w = -p \times \Delta V$ $1\text{ atm} = 1.01 \times 10^5\text{ Pa}$, and $1\text{ liter} = 10^{-3}\text{ m}^3$, so: $W = (1.01 \times 10^5\text{ N/m}^2) \times (5.35 - 1.24) \times 10^{-3}\text{ m}^3$ $= 415.11\text{J}$ So $\Delta E = -121 + 415.11$ 294 But with correct values its 291
What is the work associated with decomposition of trinitrotoluene (TNT) upon detonation according to the following reaction at 281.65 KK? $\text{C}_7\text{H}_5\text{N}_3\text{O}_6(\text{s}) \rightarrow 3\text{N}_2(\text{g}) + 5\text{H}_2(\text{g}) + 3\text{O}_2(\text{g}) + 7\text{C}(\text{s})$	1 mole $\text{C}_7\text{H}_5\text{N}_3\text{O}_6$ gives $3/2$ moles N_2 and $5/2$ moles H_2 and 3 moles O_2 gas Total gas moles produced = N_2 moles + H_2 moles + O_2 moles $= 1.5 + 2.5 + 3 = 7$ work done = $-P \times dV$ where dV is change in volume we use ideal gas equation $PV = nRT$ (since V is nothing but volume occupied all gases) work done = $-nRT$ where R = gas constant = $0.08206\text{ liter atm/molK}$ $= -7\text{ mol} \times 0.08206\text{ liter atm/molK} \times 281.65$ $= -162.8500\text{ atm liter /mol}$ we have $1\text{ atm liter} = 101.325\text{ J}$ hence work done = $-162.8500\text{ atm liter /mol} \times (101.325\text{ J/atmliter})$ $= -16500\text{ J/mol}$ $= -16.500\text{ KJ/mol}$ (-ve sign indicates expansion work) (correct answer is -16.39)

Mastering Chemistry Answers Assignment 1

Tobias Bleicker

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