

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :

[7]

- (i) The Arrhenius equation is _____.
- (a) $k = A e^{\frac{RT}{E_a}}$ (b) $A = k e^{\frac{-E_a}{RT}}$
- (c) $k = A e^{\frac{-RT}{E_a}}$ (d) $k = A e^{\frac{-E_a}{RT}}$
- (ii) If the enthalpy of vaporisation of water at 100°C is 186.5 J . mol⁻¹, the entropy of vaporization will be _____.
- (a) 4.0 J . K⁻¹. mol⁻¹ (b) 3.0 J . K⁻¹. mol⁻¹
- (c) 1.5 J . K⁻¹. mol⁻¹ (d) 0.5 J . K⁻¹. mol⁻¹
- (iii) The atomicity of sulphur in orthorhombic sulphur is _____.
- (a) 8 (b) 6 (c) 4 (d) 2
- (iv) The major binding force in diamond is _____.
- (a) Covalent bond (b) Ionic bond
- (c) Metallic bond (d) Co-ordinate covalent bond
- (v) The boiling point of water at high altitude is low, because _____.
- (a) the temperature is low. (b) the atmospheric pressure is low.
- (c) the temperature is high. (d) the atmospheric pressure is high.
- (vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is _____.
- (a) 90 mhos.cm². (b) 110 mhos.cm².mol⁻¹
- (c) 200 mhos.cm².mol⁻¹ (d) 400 mhos.cm².mol⁻¹
- (vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent?
- (a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination

Q. 2. Answer any THREE of the following :

[9]

- (i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to –
- (a) Magnetic property (b) Oxidation state (c) Hydrides
- (ii) What is the value of $\Delta S_{\text{surr.}}$ for the following reaction at 298 K –
- $$6\text{CO}_{2(g)} + 6\text{H}_2\text{O}_{(l)} \longrightarrow \text{C}_6\text{H}_{12}\text{O}_{6(s)} + 6\text{O}_{2(g)}$$
- Given that : $\Delta G^\circ = 2879 \text{ kJ mol}^{-1}$, $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$.
- (iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.
- (iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol⁻¹) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46°C. Calculate the molar mass of the unknown compound.

Hsc Chemistry Test Paper 2014

Alexander Durie



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Science and Technology Education and Future Human Needs J. L. Lewis,P. J. Kelly,2014-05-17 Science and Technology Education and Future Human Needs is a collection of papers that tackle concerns in the education of future scientists particularly concerns in identifying techniques and resource material The title first covers the impact of science on society and then proceeds to tackling the relevance of science Next the selection talks about the revision of science curricula Chapter 4 deals with science education and the needs of developing countries while Chapter 5 talks about problems in implementation The sixth chapter covers the balance between technology and environment in development and the seventh chapter tackles the nutritional concerns in national development In the last chapter the text talks about addressing human needs first before developing science and technology The book will be of great interest to individuals concerned with the progress of science and technology

Understanding Student Participation and Choice in Science and Technology Education Ellen Karoline Henriksen,Justin Dillon,Jim Ryder,2014-09-17 Drawing on data generated by the EU's Interests and Recruitment in Science IRIS project this volume examines the issue of young people's participation in science technology engineering and mathematics education With an especial focus on female participation the chapters offer analysis deploying varied theoretical frameworks including sociology social psychology and gender studies The material also includes reviews of relevant research in science education and summaries of empirical data concerning student choices in STEM disciplines in five European countries Featuring both quantitative and qualitative analyses the book makes a substantial contribution to the developing theoretical agenda in STEM education It augments available empirical data and identifies strategies in policy making that could lead to improved participation and gender balance in STEM disciplines The majority of the chapter authors are IRIS project members with additional chapters written by specially invited

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