

# STD XII - CHEMISTRY

Time : 3 hrs.]

MARCH - 2013

[Marks : 70

## SECTION - I

**Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :**

[7]

- (i) In body centred cubic structure the space occupied is about \_\_\_\_\_.  
 (a) 68% (b) 53% (c) 38% (d) 32%
- (ii) For a gaseous reaction the unit of rate of reaction is \_\_\_\_\_.  
 (a)  $\text{L atm s}^{-1}$  (b)  $\text{atm mol}^{-1} \text{s}^{-1}$  (c)  $\text{atm s}^{-1}$  (d)  $\text{mol s}$
- (iii) Which of the following compounds contains S = O as well as S = S bonds?  
 (a) Sulphuric acid (b) Thiosulphuric acid  
 (c) Sulphurous acid (d) Thiosulphurous acid
- (iv) Which of the following solutions shows maximum depression in freezing point ?  
 (a) 0.5 M  $\text{Li}_2\text{SO}_4$  (b) 1 M NaCl (c) 0.5 M  $\text{Al}_2(\text{SO}_4)_3$  (d) 0.5 M  $\text{BaCl}_2$
- (v) For a chemical reaction  $dS = -0.035 \text{ kJ/K}$  and  $dH = -20 \text{ kJ}$ . At what temperature does the reaction turn nonspontaneous?  
 (a) 5.14 K (b) 57.14 K (c) 571.4 K (d) 5714.0 K
- (vi) The standard e.m.f. of the following cell is 0.463 V  

$$\text{Cu}_{(s)} \mid \text{Cu}_{(aq)}^{++} (1\text{M}) \parallel \text{Ag}_{(aq)}^{++} (1\text{M}) \mid \text{Ag}_{(s)}$$
 . What is the standard potential of Cu electrode ?  
 (a) 1.137 V (b) 0.337 V (c) 0.463 V (d) -0.463 V
- (vii)  $\text{Fe}_2\text{O}_3$  is reduced to spongy iron near the top of blast furnace by \_\_\_\_\_.  
 (a)  $\text{H}_2$  (b) CaO (c)  $\text{SiO}_2$  (d) CO

**Q. 2. Answer any SIX of the following :**

[12]

- (i) Distinguish between crystalline solid and amorphous solid.
- (ii) State Kohlrausch Law and write mathematical expression of molar conductivity of the given solution at infinite dilution.
- (iii) Write cell reactions in lead storage battery during discharge.
- (iv) Draw structures and write geometry of  $\text{PCl}_3$  and  $\text{PCl}_5$ .
- (v) Prove that  $\Delta H = \Delta U + \Delta nRT$ . What is the condition under which  $\Delta U = \Delta H$  ?
- (vi) Mention names and formulae of two ores of aluminium.
- (vii) Derive the relationship between relative lowering of vapour pressure and molar mass of nonvolatile solute.
- (viii) What is pseudo first order reaction? Give one example of it.

**Q. 3. Answer any THREE of the following :**

[9]

- (i) Calculate the mole fraction and molality of  $\text{HNO}_3$  in a solution containing 12.2%  $\text{HNO}_3$ . (Given - atomic masses : H = 1, N = 14, O = 16)
- (ii) Consider the reaction -



What are the values of the following at the same time ?

- (a)  $-\frac{d[\text{I}^-]}{dt}$  (b)  $-\frac{d[\text{S}_2\text{O}_8^{2-}]}{dt}$  (c)  $-\frac{d[\text{I}_3^-]}{dt}$

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the

copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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