

### Abstract

## Chapter Test B

## Atomic Physics

## MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- What is the frequency of a photon with an energy of  $1.99 \times 10^{-19} \text{ J}$ ?  
( $h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$ )
  - $1.00 \times 10^{14} \text{ Hz}$
  - $2.00 \times 10^{14} \text{ Hz}$
  - $3.00 \times 10^{14} \text{ Hz}$
  - $4.00 \times 10^{14} \text{ Hz}$
- Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
  - 1 : 1
  - 2 : 1
  - 3 : 1
  - 4 : 1
- A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
  - 4.6 eV
  - 4.8 eV
  - 1.4 eV
  - 2.4 eV
- What causes the bright lines in the emission spectrum of an element to occur?
  - Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
  - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
  - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
  - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
- What causes the dark lines in the absorption spectrum of an element to occur?
  - Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
  - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
  - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
  - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

# Holt Science Test B Answers

**Rosina Ehmann**



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