

**Skill** – The student will be able to calculate the equilibrium constant, given the equilibrium concentrations which refer to the equation in level 2 and tell if the products or reactants are favored.

When a reaction is said to have reached equilibrium, the ratio of the concentration of products to reactants is equal to a value called the equilibrium constant ( $K$ ). A small value of  $K$  indicates that the concentration of products is small and the concentration of reactants is large. We say the equilibrium lies to the left because there are more reactants at equilibrium. If the equilibrium constant is large, this indicates that there is a high concentration of products and a low concentration of reactants. In this situation we say the equilibrium lies to the right because there are more products at equilibrium.

Balance the equations and use the info below to calculate the equil. constant ( $K$ ) for the following reactions. Circle whether the equilibrium would shift to the right or to the left.



**Skill** – The student will be able to use  $K_{sp}$  correctly in an equilibrium calculation.  
Write the balanced chemical equations describing the dissolving of the following solids in water. Write the  $K_{sp}$  expression for each.



Use your knowledge of the relationship between  $K_{sp}$ , molarity, & solubility ( $\text{g/L}$ ) to solve the following:



# Equilibrium Unit Webquest And Notes Answers

**AW Rasmussen**



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