

STO.1	Balance a chemical equation.
STO.2	Identify the parts of a chemical equation.
RXN.1	Describe a chemical reaction using words and symbolic equations.

For each of the following problems, write complete chemical equations to describe the chemical process taking place. Balance the equations.

- 1) When lithium hydroxide pellets are added to a solution of sulfuric acid ( $\text{H}_2\text{SO}_4$ ), lithium sulfate and water are formed.
- 2) Magnesium reacts with sodium fluoride to produce magnesium fluoride and elemental sodium.
- 3) If a copper coil is placed into a solution of silver nitrate, silver crystals form and copper (I) nitrate is generated.
- 4) When crystalline  $\text{C}_6\text{H}_{12}\text{O}_6$  is burned in oxygen, carbon dioxide and water vapor are formed.
- 5) Calcium carbonate combines with hydrochloric acid ( $\text{HCl}$ ) to produce calcium chloride, water and carbon dioxide gas.

Balance the equations below:

- 1) \_\_\_\_\_  $\text{N}_2$  + \_\_\_\_\_  $\text{H}_2 \rightarrow$  \_\_\_\_\_  $\text{NH}_3$
- 2) \_\_\_\_\_  $\text{KClO}_3 \rightarrow$  \_\_\_\_\_  $\text{KCl}$  + \_\_\_\_\_  $\text{O}_2$
- 3) \_\_\_\_\_  $\text{NaCl}$  + \_\_\_\_\_  $\text{F}_2 \rightarrow$  \_\_\_\_\_  $\text{NaF}$  + \_\_\_\_\_  $\text{Cl}_2$
- 4) \_\_\_\_\_  $\text{H}_2$  + \_\_\_\_\_  $\text{O}_2 \rightarrow$  \_\_\_\_\_  $\text{H}_2\text{O}$
- 5) \_\_\_\_\_  $\text{Pb}(\text{OH})_2$  + \_\_\_\_\_  $\text{HCl} \rightarrow$  \_\_\_\_\_  $\text{H}_2\text{O}$  + \_\_\_\_\_  $\text{PbCl}_2$
- 6) \_\_\_\_\_  $\text{AlBr}_3$  + \_\_\_\_\_  $\text{K}_2\text{SO}_4 \rightarrow$  \_\_\_\_\_  $\text{KBr}$  + \_\_\_\_\_  $\text{Al}_2(\text{SO}_4)_3$
- 7) \_\_\_\_\_  $\text{CH}_4$  + \_\_\_\_\_  $\text{O}_2 \rightarrow$  \_\_\_\_\_  $\text{CO}_2$  + \_\_\_\_\_  $\text{H}_2\text{O}$
- 8) \_\_\_\_\_  $\text{C}_3\text{H}_8$  + \_\_\_\_\_  $\text{O}_2 \rightarrow$  \_\_\_\_\_  $\text{CO}_2$  + \_\_\_\_\_  $\text{H}_2\text{O}$
- 9) \_\_\_\_\_  $\text{C}_6\text{H}_{18}$  + \_\_\_\_\_  $\text{O}_2 \rightarrow$  \_\_\_\_\_  $\text{CO}_2$  + \_\_\_\_\_  $\text{H}_2\text{O}$
- 10) \_\_\_\_\_  $\text{FeCl}_3$  + \_\_\_\_\_  $\text{NaOH} \rightarrow$  \_\_\_\_\_  $\text{Fe}(\text{OH})_3$  + \_\_\_\_\_  $\text{NaCl}$
- 11) \_\_\_\_\_  $\text{P}$  + \_\_\_\_\_  $\text{O}_2 \rightarrow$  \_\_\_\_\_  $\text{P}_2\text{O}_5$

# Equation Writing And Balancing 8

**F Rizvi**



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