

Equation Sheet for A-Level Chemistry

These are the equations you need to

memorise for your Chemistry A-Level



Amount of Substance

Moles
 $n = \frac{m}{M_r}$

Concentration
 $C = \frac{n}{V}$

Molar gas volume
 $V = n \times 24$

Number of particles
 $= n \times N_A$

Ideal gas Equation
 $pV = nRT$

Atom economy
 $= \frac{\text{M}_r \text{ of desired product}}{\text{Sum of M}_r \text{ of all reactants}} \times 100$

% Yield
 $= \frac{\text{Actual yield}}{\text{Theoretical yield}} \times 100$

Energetics

Enthalpy change
 $Q = mc\Delta T$

Average bond enthalpy
 $\Delta H = \Sigma (\text{bond enthalpies in reactants}) - \Sigma (\text{bond enthalpies in products})$

Electrode Potentials

EMF
 $E^\ominus = E^\ominus \text{right hand side} - E^\ominus \text{left hand side}$

Thermodynamics

Gibbs Free Energy Equation
 $\Delta G = \Delta H - T\Delta S$

Transition Metals

Coloured complex ions - difference in energy
 $\Delta E = h\nu = \frac{hc}{\lambda}$

Chemical Equilibria

Equilibrium constant wrt concentration
 $K_c = \frac{[D]^d [E]^e}{[A]^a [B]^b}$

Equilibrium constant wrt pressure
 $K_p = \frac{(p_D)^d (p_E)^e}{(p_A)^a (p_B)^b}$

Rate Equations

Rate = $k[A]^m[B]^n$

Arrhenius is provided but you should be able to use:

$k = Ae^{-\frac{E_a}{RT}}$
 $\ln k = -E_a/RT + \ln A$

Acids and Bases

$pH = -\log_{10} [H^+]$

$K_w = [H^+][OH^-]$

$pK_a = -\log_{10} K_a$

$K_a = \frac{[H^+][A^-]}{[HA]}$

These are the meaning of the symbols on the equation sheet



Amount of Substance

n = number of moles
 m = mass
 M_r = formula mass
 C = concentration
 V = volume
 N_A = Avogadro's constant
 p = pressure
 R = gas constant
 T = temperature

Rate Equations

k = rate constant
 $[A]$ = concentration of A
 $[B]$ = concentration of B
 a = order wrt A
 b = order wrt B
 A = Arrhenius pre exponential factor
 e = exponential
 E_a = activation energy

Energetics

q = heat energy
 m = mass
 c = specific heat capacity
 ΔT = change in temperature
 ΔH = enthalpy change

Electrode Potentials

$E^\ominus$ = electrode potential

Thermodynamics

ΔG = Gibbs free energy
 ΔH = enthalpy change
 T = temperature
 ΔS = change in entropy

Transition Metals

ΔE = change in energy
 h = Planck's constant
 ν = frequency

Chemical Equilibria

K_c = Equilibrium constant (wrt concentration)
 K_p = Equilibrium constant (wrt pressure)
 $[A] / (A)$ = concentration/ partial pressure of reactant A
 $[B] / (B)$ = concentration/ partial pressure of reactant B
 $[D] / (D)$ = concentration/ partial pressure of product D
 $[E] / (E)$ = concentration/ partial pressure of product E
 a = moles of A from balanced equation
 b = moles of B from balanced equation
 d = moles of D from balanced equation
 e = moles of E from balanced equation

Acids and Bases

$[H^+]$ = concentration of H^+
 $[OH^-]$ = concentration of OH^-
 $[A^-]$ = concentration of salt

Mcmuc A Level Chemistry

**Horace George Andrew, J. G. D.
Carpenter, David Abbott**



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