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ACID-BASE EQUILIBRIA

Arrhenius Acid:

Anything that produces hydrogen ions in aqueous solution

Arrhenius Base:

Anything that produces hydroxide ions in aqueous solution

Brønsted-Lowry Acid:

A species that donates a proton is an acid

Brønsted-Lowry Base:

A species that accepts a proton is a base

Amphiprotic Species: a chemical species that can both donate and accept protons

Lewis Acid:

Anything that accepts a pair of electrons

Lewis Base:

Anything that donates a pair of electrons

Water Equilibrium Constant (K_w):

$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.00 \times 10^{-14}$ (at 25°C)

Strong Acids:

Acids that donate protons to water molecules quantitatively

Strong Base:

A substance that generates hydroxide ions quantitatively in aqueous solution

pH SCALES:

$\text{pH} = -\log[\text{H}_3\text{O}^+]$ $\text{pOH} = -\log[\text{OH}^-]$ $\text{p}K_a = -\log K_a$

$\text{p}K_b = -\log K_b$ $\text{pH} + \text{pOH} = 14.00$ $K_a K_b = K_w$ $\text{p}K_a + \text{p}K_b = 14.00$

Weak Acid:

Acid that reaches equilibrium when only a small fraction of its molecules transfer protons to water

Weak Base:

Generates hydroxide ions by accepting protons from water but reaches equilibrium when only a fraction of its molecules have done so

APPLICATION OF Aqueous Equilibria:

BUFFER SOLUTION:

Contains both a weak acid and its conjugate base as major species in solution

$$\text{pH} = \text{p}K_a + \log \left(\frac{[\text{A}^-]_{\text{total}}}{[\text{HA}]_{\text{total}}} \right)$$

BUFFER CAPACITY:

The amount of added H_3O^+ or OH^- the buffer solution can tolerate without exceeding a specified pH range

CHEMICAL KINETICS

REACTION MECHANISM:

The exact molecular pathway that starting materials follow on their way to becoming products

RATE-DETERMINING STEP:

The slowest elementary step in a mechanism

FIRST-ORDER RATE LAW:

Rate = $k[A]$, where A is a reactant in the overall reaction, this can be converted to

$$\ln \left(\frac{[A]_0}{[A]} \right) = kt$$

HALF-LIFE ($t_{1/2}$), 1st ORDER:

When half the original concentration has been consumed,

$$[A] = 0.5[A]_0 \quad t_{1/2} = \frac{\ln 2}{k}$$

SECOND-ORDER RATE EXPRESSION:

Rate = $k[A]^2$, this can be converted to

$$\frac{1}{[A]} - \frac{1}{[A]_0} = kt$$

BI-MOLECULAR ELEMENTARY REACTION:

$A + B \rightarrow \text{products}$ Elementary rate = $k[A][B]$

UNI-MOLECULAR ELEMENTARY REACTION:

$C \rightarrow \text{products}$ Elementary rate = $k[C]$

FIRST STEP IS RATE-DETERMINING:

The predicted rate law for the overall reaction is the rate expression for that first step

ACTIVATION ENERGY (E_a):

Energy barrier, the minimum energy that must be supplied before the reaction can occur

Activated Complex: the molecular arrangement at the point of highest energy along the energy level diagram

CHEMICAL EQUILIBRIUM

EQUILIBRIUM CONSTANT:

$$K_{eq} = \frac{[\text{D}]_{eq}^d [\text{E}]_{eq}^e}{[\text{A}]_{eq}^a [\text{B}]_{eq}^b}$$

K_{eq} applied only at equilibrium, K_{eq} is independent of initial conditions, K_{eq} is related to the stoichiometry

Pure Liquid and Solid:

The concentrations of pure liquids or solids are always equal to their standard concentrations, therefore division by standard concentration results in a value of 1

LARGE K_{eq} :

Indicates that the reaction goes virtually to completion

Le Chatelier's Principle:

When a change is imposed on a system at equilibrium, the system will react in the direction that reduces the amount of change

TEMPERATURE:

The only variable that causes a change in the value of K_{eq} , an increase in temperature always shifts the equilibrium position in the endothermic direction

THERMODYNAMICS

CLOSED SYSTEM:

Exchanges energy but not matter with its surroundings

ISOLATED SYSTEM:

Exchanges neither matter nor energy with the surroundings

STATE VARIABLES:

Conditions that must be specified to establish the state of a system, pressure (P), volume (V), temperature (T), and amounts of substances (n)

PHYSICAL CHANGE OF STATE:

Some of the state variables changes, but the chemical composition of the system stays the same

CHEMICAL CHANGE OF STATE:

The amounts of reactants and products change

Chemical Reaction: some bonds break, and new bonds form

STATE FUNCTION:

A property that depends only on the state of the system

PATH FUNCTION:

A property that depends on how a change takes place

Molar Heat Capacity:

The amount of heat needed to raise the temperature of 1 mol of substance by 1 Kelvin (1K)

WORK (W):

Energy used to move an object against an opposing force, $w = Fd$

FIRST LAW OF THERMODYNAMICS:

$$\Delta E_{\text{sys}} = q_{\text{sys}} + w_{\text{sys}}$$

Exothermic Process:

If the chemicals release heat, this heat gain raises the temperature of the surroundings

Endothermic Process:

If the chemicals absorb heat, this heat loss lowers the temperature of the surroundings

$$q_{\text{surroundings}} = -C_{\text{cal}} \Delta T, \quad q = nC_p \Delta T$$

ENTHALPY:

A thermodynamic quantity whose change equals the heat flow at constant pressure

$$\Delta H = \Delta E + \Delta(PV)$$

STANDARD ENTHALPY OF FORMATION (ΔH_f°):

Enthalpy change accompanying the formation of one mole of a chemical substance from pure elements in their most stable forms under standard conditions

Hess' Law:

The enthalpy change for any overall process is equal to the sum of enthalpy changes for any set of steps that leads from the reactants to the products

MOLEAR HEAT OF SOLUTION:

Measures net energy flow that occurs as substance dissolves

MOLEAR HEAT OF VAPORIZATION:

The heat needed to vaporize one mole of a substance at its normal boiling point

MOLEAR HEAT OF FUSION:

Heat needed to melt one mole of a substance at its normal melting point

SUBLIMATION:

A phase change in which a solid converts directly to a vapour without passing through the liquid phase

$$\Delta E_{\text{sys}} = \Delta H_{\text{vap}} - RT_{\text{vap}}$$

SECOND LAW OF THERMODYNAMICS:

Any spontaneous process increases the disorder of the universe

ENTROPY:

The state function that provides a quantitative measure of disorder and is symbolized S,

$$\Delta S = \frac{q_r}{T}$$

Entropy Change of the Universe:

Total entropy change,

$$\Delta S_{\text{universe}} = \Delta S_{\text{system}} + \Delta S_{\text{surroundings}}$$

REACTION ENTROPIES:

$$\Delta S_{\text{reaction}} = \sum \text{coeff}_p \Delta S^\circ - \sum \text{coeff}_r \Delta S^\circ$$



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