

- Q.65** The nitrogenous base different in RNA as compared to DNA is  
 A) Cytosine  
 B) Adenine  
 C) Thymine  
 D) Guanine
- Q.66** Carbohydrates are stored in human body as the polysaccharide  
 A) Glycogen  
 B) Starch  
 C) Glucose  
 D) Galactase
- Q.67** Denaturation of protein is used by  
 A) Addition of sulphuric acid  
 B) Addition of dichromate ion  
 C) Changing the pH  
 D) All of these
- Q.68** In the reaction,  $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$ , when one mole of ammonia and one mole of oxygen are made to react to completion, then  
 A) 1.0 mol of  $\text{H}_2\text{O}$  is produced  
 B) 1.5 mol of NO is formed  
 C) All the oxygen is consumed  
 D) All the ammonia is consumed
- Q.69** In which set all isotopes have maximum natural abundance  
 A)  $^{40}_{20}\text{Ca}$ ,  $^{12}_6\text{C}$ ,  $^{35}_{17}\text{Cl}$   
 B)  $^{40}_{20}\text{Ca}$ ,  $^{56}_{26}\text{Fe}$ ,  $^{24}_{12}\text{Mg}$   
 C)  $^{40}_{20}\text{Ca}$ ,  $^{23}_{11}\text{Na}$ ,  $^{37}_{17}\text{Cl}$   
 D)  $^{20}_{10}\text{Ne}$ ,  $^{32}_{16}\text{S}$ ,  $^{14}_7\text{N}$
- Q.70** Two balloons are filled with equal moles of hydrogen and helium gas at STP. Which balloon will contract first if same size holes are made in both of them  
 A) Both will contract at same time  
 B) Balloon filled with hydrogen gas  
 C) Balloon filled with Helium gas  
 D) Only hydrogen filled balloon will contract
- Q.71** In all living organisms, nucleic acids are responsible to their ability  
 A) To reproduce and store genetic information  
 B) To transmit genetic information  
 C) To undergo mutation  
 D) All of these
- Q.72** Ozone acts as pollutant when present in  
 A) Troposphere  
 B) Stratosphere  
 C) Mesosphere  
 D) Thermosphere
- Q.73** A primary pollutant is that  
 A) Which causes cancer  
 B) Is waste product of an industry  
 C) Whose concentration cannot be controlled  
 D) Which corrodes the metals only
- Q.74** The pH of a neutral sample of water at  $100^\circ\text{C}$  when  $K_w = 1.0 \times 10^{-12}$  is  
 A) 7  
 B) 6  
 C) 0  
 D) 12
- Q.75** The two steps involved in the gas phase reaction  $\text{X} + 2\text{Y} \rightarrow \text{XY}_2$  are shown below  
 $\text{X} + \text{Y} \xrightarrow{\text{slow}} \text{XY}$ ,  $\text{XY} + \text{Y} \xrightarrow{\text{fast}} \text{XY}_2$   
 What is the rate equation for the overall reaction  
 A) Rate =  $[\text{x}]^1[\text{y}]^1$   
 B) Rate =  $[\text{x}]^1[\text{y}]^2$   
 C) Rate =  $[\text{x}]^0[\text{y}]^2$   
 D) Rate =  $[\text{xy}]^1[\text{y}]^1$
- Q.76** Which of the following statements about the reaction given below is/are correct  
 $\text{IO}_3^-(\text{aq}) + 2\text{I}^-(\text{aq}) + 6\text{H}^+(\text{aq}) + 6\text{Cl}^-(\text{aq}) \rightarrow 3\text{ICl}_2(\text{aq}) + 3\text{H}_2\text{O}(\text{l})$   
 A) The oxidation number of iodine in the iodate ion,  $\text{IO}_3^-(\text{aq})$ , changes from +5 to +1  
 B) The oxidation number of iodine in the iodide ion  $\text{I}^-$ , changes from -1 to +2  
 C) The oxidation number of chlorine in the iodide ion  $\text{I}^-$ , changes from -1 to -2  
 D) Iodate ions are oxidized
- Q.77** The one having highest ionization energy is  
 A) Mg  
 B) Al  
 C) S  
 D) P
- Q.78** The electronic configuration of elements across each period in the periodic table starts with a s-subshell of the new shell and ends with  
 A) Small subshell  
 B) The next higher shell  
 C) Different subshell of the same subshell  
 D) p-subshell of the same shell
- Q.79** Why binary compounds of oxygen and fluorine are called fluorides rather than oxides  
 A) They always contain F-ions  
 B) O atom is larger than F atom  
 C) F is more electronegative than O  
 D) O is better oxidizing agent
- Q.80** All the elements of group IVA do not have which of the property in common  
 A) Show four valency  
 B) Form tetrahalide  
 C) Show inert pair effect  
 D) Form dioxide

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