

ENGAGE

Why does the atomic number but not the mass number change in positron emission?

SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission**TRY IT FIRST**

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

SOLUTION GUIDE

	Given	Need	Connect
ANALYZE THE PROBLEM	Mn-49, positron emission	balanced nuclear equation	mass number, atomic number of new nucleus

STEP 1 Write the incomplete nuclear equation.



STEP 2 Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = ? + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49 \text{ (mass number of new nucleus)}$$

STEP 3 Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = ? + 1$$

$$25 - 1 = ?$$

$$25 - 1 = 24 \text{ (atomic number of new nucleus)}$$

STEP 4 Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written ${}^{49}_{24}\text{Cr}$.

STEP 5 Complete the nuclear equation.

**STUDY CHECK 5.4**

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

ANSWER**TEST**

Try Practice Problems 5.15 to 5.18

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Table of Contents Mastering Chemistry Answers

1. Understanding the eBook Mastering Chemistry Answers
 - The Rise of Digital Reading Mastering Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Answers
 - Personalized Recommendations
 - Mastering Chemistry Answers User Reviews and Ratings
 - Mastering Chemistry Answers and Bestseller Lists
5. Accessing Mastering Chemistry Answers Free and Paid eBooks

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 - Mastering Chemistry Answers Budget-Friendly Options
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 - Mastering Chemistry Answers Compatibility with Devices
 - Mastering Chemistry Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answers
 - Highlighting and Note-Taking Mastering Chemistry Answers
 - Interactive Elements Mastering Chemistry Answers
 8. Staying Engaged with Mastering Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Answers
 9. Balancing eBooks and Physical Books Mastering Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mastering Chemistry Answers
 - Setting Reading Goals Mastering Chemistry Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mastering Chemistry Answers
 - Fact-Checking eBook Content of Mastering Chemistry Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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