

Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point. Express the heat in kilojoules to three significant figures.

$$q = 5830 \text{ kJ}$$

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Correct

At its boiling point, the enthalpy of vaporization (ΔH_{vap}) of water is 40.7 kJ/mol. Once the given mass, $2.58 \times 10^3 \text{ g}$, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by ΔH_{vap} yields the amount of heat required to evaporate that amount of water.

Part B

Suppose that 0.48 g of water at 25°C condenses on the surface of a 55-g block of aluminum that is initially at 25°C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum, $C_{s,\text{Al}}$, is 0.903 J/(g · °C).)

Express the temperature in degrees Celsius to two significant figures.

$$T_f =$$

°C

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**T. W. Graham Solomons, Craig B.
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Table of Contents Mastering Chemistry Answer Key Chapter 12

1. Understanding the eBook Mastering Chemistry Answer Key Chapter 12
 - The Rise of Digital Reading Mastering Chemistry Answer Key Chapter 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Answer Key Chapter 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answer Key Chapter 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Answer Key Chapter 12
 - Personalized Recommendations
 - Mastering Chemistry Answer Key Chapter 12 User Reviews and Ratings
 - Mastering Chemistry Answer Key Chapter 12 and Bestseller Lists

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 6. Navigating Mastering Chemistry Answer Key Chapter 12 eBook Formats
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 - Mastering Chemistry Answer Key Chapter 12 Compatibility with Devices
 - Mastering Chemistry Answer Key Chapter 12 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answer Key Chapter 12
 - Highlighting and Note-Taking Mastering Chemistry Answer Key Chapter 12
 - Interactive Elements Mastering Chemistry Answer Key Chapter 12
 8. Staying Engaged with Mastering Chemistry Answer Key Chapter 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Answer Key Chapter 12
 9. Balancing eBooks and Physical Books Mastering Chemistry Answer Key Chapter 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answer Key Chapter 12
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mastering Chemistry Answer Key Chapter 12
 - Setting Reading Goals Mastering Chemistry Answer Key Chapter 12
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mastering Chemistry Answer Key Chapter 12
 - Fact-Checking eBook Content of Mastering Chemistry Answer Key Chapter 12
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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