

CHEM 113 Final – Monday May 16th 10:15-12:15 in LSC 133; bring a pencil!

Atomic Structure:

1. What do these have in common?



- A) The same number of protons
- B) The same number of electrons
- C) The same number of neutrons
- D) The same size

2. The element X occurs natural to the extent of 20.0% ^{12}X and 80.0% ^{13}X . The atomic mass of X is nearest

- A) 12.2
- B) 12.5
- C) 12.8
- D) 13.0

3. What is the valence electron configuration for the element in Period 5, Group 3A?

- A) $5s^25p^1$
- B) $3s^23p^3$
- C) $3s^23p^5$
- D) $5s^25p^3$

4. What is the maximum number of electrons that can occupy an orbital labeled d_{xy} ?

- A) 1
- B) 2
- C) 4
- D) 10

5. Which electron transition in a hydrogen atom is associated with the largest emission of energy?

- A) $n = 2$ to $n = 1$
- B) $n = 1$ to $n = 2$
- C) $n = 2$ to $n = 4$
- D) $n = 4$ to $n = 2$

Molecular Structure and Bonding:

1. Which molecule has exactly two unshared (lone) pairs of electrons on the central atom?

- A) BF_3
- B) OF_2
- C) NF_3
- D) XeF_2

2. The molecule of the type ML_4 consists of four single bonds and no lone pairs. What structure is it expected to assume?

- A) square planar
- B) trigonal pyramidal
- C) trigonal planar
- D) tetrahedral

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Table of Contents Mastering Chemistry Chem 113 Exercise Answers

1. Understanding the eBook Mastering Chemistry Chem 113 Exercise Answers
 - The Rise of Digital Reading Mastering Chemistry Chem 113 Exercise Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Chem 113 Exercise Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Chem 113 Exercise Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Chem 113 Exercise Answers
 - Personalized Recommendations
 - Mastering Chemistry Chem 113 Exercise Answers User Reviews and Ratings
 - Mastering Chemistry Chem 113 Exercise Answers and Bestseller Lists
5. Accessing Mastering Chemistry Chem 113 Exercise Answers Free and Paid eBooks
 - Mastering Chemistry Chem 113 Exercise Answers Public Domain eBooks
 - Mastering Chemistry Chem 113 Exercise Answers eBook Subscription Services
 - Mastering Chemistry Chem 113 Exercise Answers Budget-Friendly Options
6. Navigating Mastering Chemistry Chem 113 Exercise Answers eBook Formats

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 - Mastering Chemistry Chem 113 Exercise Answers Compatibility with Devices
 - Mastering Chemistry Chem 113 Exercise Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Mastering Chemistry Chem 113 Exercise Answers
 - Highlighting and Note-Taking Mastering Chemistry Chem 113 Exercise Answers
 - Interactive Elements Mastering Chemistry Chem 113 Exercise Answers
8. Staying Engaged with Mastering Chemistry Chem 113 Exercise Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Chem 113 Exercise Answers
9. Balancing eBooks and Physical Books Mastering Chemistry Chem 113 Exercise Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Chem 113 Exercise Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mastering Chemistry Chem 113 Exercise Answers
- Setting Reading Goals Mastering Chemistry Chem 113 Exercise Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Chemistry Chem 113 Exercise Answers
- Fact-Checking eBook Content of Mastering Chemistry Chem 113 Exercise Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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