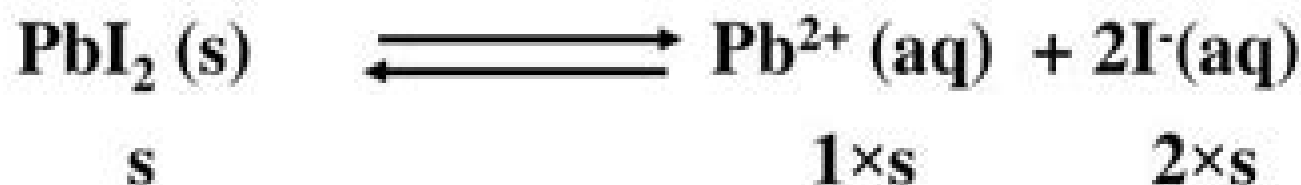


**Calculate the molar solubility of lead iodide  $\text{PbI}_2$ ,  
from its  $K_{sp}$  in water at  $25^\circ\text{C}$**



$$K_{sp} = [\text{Pb}^{2+}]^1 [\text{I}^{-}]^2$$

$$K_{sp} = [1 \times s]^1 [2 \times s]^2$$

$$K_{sp} = (1s) \times (2s) \times (2s)$$

$$K_{sp} = (1 \times 2 \times 2) \times (s \times s \times s)$$

$$K_{sp} = 4s^3$$

# Ksp Of Lead Iodide Lab Calculations

**J Elliott**



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