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- ### 3. 研究設計

☞ কোলাস পানিসূত্র লবণের দ্রাব্যতা আপনাত্মা কৃষ্ণির সাথে সাথে এক পর্যায়ে কৃষ্ণি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা কৃষ্ণির সাথে সাথে দ্রাব্যতা কৃষ্ণি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কোলাস পানি দ্রুত লবণ নিষ্কৃতিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলার যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত হওয়ার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু 32.4°C আপনাত্মার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C আপনাত্মার উপরে আপনাত্মা কৃষ্ণিজনিত কারণে দ্রাব্যতা হ্রাস ঘটে।

$w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

$$V = \frac{250}{1000} = 0.25 \text{ L}$$
$$\text{例 7-6-A: } \text{H}_2\text{SO}_4 \text{ 的 物质的量} = \left(0.5 \times \frac{2.50}{1000} \right) \text{ mol} \\ = 0.125 \text{ mol}$$
$$\begin{aligned} \text{NaOH-H: Na}_2\text{CO}_3 \text{ से द्रव्य के भार} &= \left(0.11 \times \frac{700}{1000} \right) \text{ mol} \\ &= 0.077 \text{ mol} \end{aligned}$$

अण्विहित समीकरण: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
अधिकतम घट पादि, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
0.077 mol " " " " (0.077) mol "

অনশিট H_2SO_4 এর পরিমাণ = $(0.125 - 0.077) \text{ mol}$
 = 0.048 mol

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে $0.1 \text{ mol H}_2\text{SO}_4$ অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লিক।

উদাহরণ ২ একজন ছাত্র ল্যাবরেটরিতে X-সম্পত্তিতে 10 mg এর তরুণ পরিমাণ এবং Y-সম্পত্তিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর দ্রবণ 5 মিলি ব্যবহার করে।

Abstract

- ক. নোট কাকে বলে?
- খ. Cl_2 অপেক্ষার— ব্যাখ্যা করো।
- গ. উদ্ভীপকের এসিডটির মৌল্যত্র ঘনমাত্রা নির্ণয় করো।
- ঘ. উদ্ভীপকের কোন পদার্থটি পরিবেশবান্ধব? বিশ্লেষণ করো।

२. नर आश्रित किशोर

উদাহরণ ১: নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

১৮. সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl_1 এর তড়িৎ ঋণাত্মকতা 3.0 ও Cl_2 অপূর্ণ তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_1 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উদাহরণ মৌলিক অণুর ক্ষেত্রে এদৃশ ঘটার

গ) উদ্দীপকে বর্ণিত এসিডটি হলো 2% $\left(\frac{W}{V}\right)$ HCl।

2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—

100 mL. HCl এর উষ্ণ 2g

$$\therefore 1000 \text{ mL} \times \frac{2 \times 1000}{100} \text{ g} = 20 \text{ g}$$

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Minjie Lin



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