

Experiment 10

Oxidation-Reduction (Redox) Reactions

Pre-Lab Assignment

Before coming to lab:

- Read the lab thoroughly.
- Answer the pre-lab questions that appear at the end of this lab exercise.

Purpose

A set of oxidation-reduction (redox) reactions will be performed to determine the relative strengths as reducing and oxidizing agents of metals and other ions. An activation series will be devised and used to predict whether a reaction will occur spontaneously or not.

Background

Oxidation-reduction (redox) reactions are a classification of chemical changes that involve the transfer of electrons. An example of a redox reaction is shown in Eqn. 1, when magnesium metal reacts with chlorine gas.



At first glance, the electron exchange is not apparent as electrons are not included in the overall balanced equation. However, electrons are shown when the reaction is split into halves: an oxidized half-reaction and a reduced half-reaction, as in Eqn. 2 and Eqn. 3 respectively.



Notice that in the half-reactions, the electrons are required to balance the charges. Since Mg is becoming more positive, one magnesium atom loses two electrons to become a cation and is said to be **oxidized**. Chlorine is becoming more negative, so one molecule of $\text{Cl}_2(\text{g})$ gains two electrons to become two chloride anions and is said to be **reduced**. Both oxidation and reduction happen simultaneously in a single redox reaction. Overall, Mg is transferring two electrons to chlorine (Fig. 1).

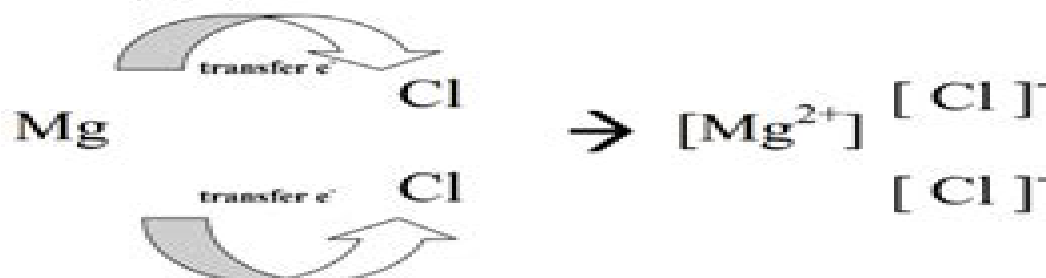


Fig. 1: Redox reaction between Mg(s) and Cl₂(g) to produce MgCl₂(aq)

Lab Manual Answers Redox Oxidation Reduction Reaction

Anju Rawlley,D. V. Saraf



Lab Manual Answers Redox Oxidation Reduction Reaction:

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 *Laboratory Manual for Principles of General Chemistry* 11th Edition covers two semesters of a general chemistry laboratory program. The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure, selecting the proper apparatus, employing the proper techniques, systematically analyzing and interpreting the data, and minimizing inherent variables. As a result of good data, a scientific and analytical conclusion is made, which may or may not be right but is certainly consistent with the data. Experiments write textbooks; textbooks don't write experiments. A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered. Further experimentation provides additional cause-effect observations leading to an even better understanding of the experiment. The 11th edition's experiments are informative and challenging while offering a solid foundation for technique, safety, and experimental procedure. The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment. The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit. An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment. Continued emphasis on handling data is supported by the Data Analysis section. Environmental Sampling and Analysis Maria Csuros, 1997-03-24. This manual covers the latest laboratory techniques, state of the art instrumentation, laboratory safety, and quality assurance and quality control requirements. In addition to complete coverage of laboratory techniques, it also provides an introduction to the inorganic nonmetallic constituents in environmental samples, their chemistry, and their control by regulations and standards. *Environmental Sampling and Analysis Laboratory Manual* is perfect for college and graduate students learning laboratory practices as well as consultants and regulators who make evaluations and quality control decisions. Anyone performing laboratory procedures in an environmental lab will appreciate this unique and valuable text.

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Study Guide and Solutions Manual Neil E. Schore,2002-08-02

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