

The University of the State of New York  
REGENTS HIGH SCHOOL EXAMINATION**ALGEBRA 2/TRIGONOMETRY**

Wednesday, August 18, 2010—8:30 to 11:30 a.m., only

Student Name: Mr. SibolSchool Name: HS For Civil Rights

Print your name and the name of your school on the lines above. Then turn to the last page of this booklet, which is the answer sheet for Part I. Fold the last page along the perforations and, slowly and carefully, tear off the answer sheet. Then fill in the heading of your answer sheet.

This examination has four parts, with a total of 39 questions. You must answer all questions in this examination. Write your answers to the Part I multiple-choice questions on the separate answer sheet. Write your answers to the questions in Parts II, III, and IV directly in this booklet. All work should be written in pen, except graphs and drawings, which should be done in pencil. Clearly indicate the necessary steps, including appropriate formula substitutions, diagrams, graphs, charts, etc.

The formulas that you may need to answer some questions in this examination are found at the end of the examination. This sheet is perforated so you may remove it from this booklet.

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When you have completed the examination, you must sign the statement printed at the end of the answer sheet, indicating that you had no unlawful knowledge of the questions or answers prior to the examination and that you have neither given nor received assistance in answering any of the questions during the examination. Your answer sheet cannot be accepted if you fail to sign this declaration.

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A graphing calculator and a straightedge (ruler) must be available for you to use while taking this examination.

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# June 2008 Global Jmap Answer Key

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web elements in the second row of the periodic table place their electrons in the  $2n$  shell as well as the  $1n$  shell for instance lithium li has three electrons two fill the  $1s$  orbital and the third is placed in the  $2s$  orbital giving an electron configuration of  $1s^2 2s^1$

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web the electron removed during the ionization of beryllium he 2s 2 is an s electron whereas the electron removed during the ionization of boron he 2s 2 2p 1 is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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