

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 \text{ kPa} / 1 \text{ atm}$ to get
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$\text{or } P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$\text{or } P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

Ideal Gas Answers Show Work

F Rizvi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Ideal Gas Answers Show Work

1. Understanding the eBook Ideal Gas Answers Show Work
 - The Rise of Digital Reading Ideal Gas Answers Show Work
 - Advantages of eBooks Over Traditional Books
2. Identifying Ideal Gas Answers Show Work
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ideal Gas Answers Show Work
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ideal Gas Answers Show Work
 - Personalized Recommendations
 - Ideal Gas Answers Show Work User Reviews and Ratings
 - Ideal Gas Answers Show Work and Bestseller Lists
5. Accessing Ideal Gas Answers Show Work Free and Paid eBooks
 - Ideal Gas Answers Show Work Public Domain eBooks
 - Ideal Gas Answers Show Work eBook Subscription Services
 - Ideal Gas Answers Show Work Budget-Friendly Options
6. Navigating Ideal Gas Answers Show Work eBook Formats

- ePub, PDF, MOBI, and More
- Ideal Gas Answers Show Work Compatibility with Devices
- Ideal Gas Answers Show Work Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ideal Gas Answers Show Work
 - Highlighting and Note-Taking Ideal Gas Answers Show Work
 - Interactive Elements Ideal Gas Answers Show Work
- 8. Staying Engaged with Ideal Gas Answers Show Work
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ideal Gas Answers Show Work
- 9. Balancing eBooks and Physical Books Ideal Gas Answers Show Work
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ideal Gas Answers Show Work
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ideal Gas Answers Show Work
 - Setting Reading Goals Ideal Gas Answers Show Work
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ideal Gas Answers Show Work
 - Fact-Checking eBook Content of Ideal Gas Answers Show Work
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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component labeling this is a computer vision algorithm to detect and count the number of connected regions also called blobs in a binary image often this is done after a segmentation algorithm

connected component labelling through matlab github - Jun 07 2022

web may 16 2019 this is a matlab function that utilises the image processing toolbox to convert any image to its edge image using canny algorithm and then label connected components in it the whole labelling algorithm is in a function that you can apply to any image the edge image goes through the labelling algorithm and final image is displayed

connected component labeling file exchange matlab - Mar 16 2023

web feb 10 2014 this function is partly based on the two pass algorithm described here en wikipedia org wiki connected component labeling i did not make use of the disjoint data set structure at least that i am aware of to connect the linked labels