

## The Hydrogen Atom – Student Guide

### Background Material

Carefully read the background pages created through Eureka, Slides, and Document and answer the following questions in detail and understanding.

Question 1: Complete the following table which compares how the Bohr model and the Quantum model compare the Hydrogen atom. In each case, they both describe things in a unique way and in some ways they do not.

| Concept<br>Name                              | Bohr Model                               | Quantum Model                            |
|----------------------------------------------|------------------------------------------|------------------------------------------|
| Electrons                                    | positively charged particles in the atom | negatively charged particles in the atom |
| Orbitals                                     | negatively charged and orbit the nucleus | negatively charged clouds                |
| Location of the electron in the atom without | is in two fixed paths                    | electron cloud particles                 |
| Energy of the electron in the atom without   | always has the same energy               | varies the energy level                  |
| Energy of the electron in the atom without   | $E = -13.6 \text{ eV}$                   | $E = -13.6 \text{ eV}$                   |

Question 2: Complete the following table which shows the comparison of two different theories by putting predictions made by each theory side-by-side. The table can be completed by using the Bohr model or Quantum model. Significant differences, similar responses, similar energy, and the same results through calculations are all fine.

# Hydrogen Atom Student Guide Naap Answers

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*Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry* Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

**The Hydrogen Atom Energy Levels**, 1976 **Student's Guide to Chemistry; a Modern Introduction** David Brooks, 1974 The Reaction of Hydrogen Atom with Hydrogen Molecular Ion Dion Allan Stams, 1985 **On the Energy Levels of a Model of the Compressed Hydrogen Atom** S. R. de Groot, C. A. ten Seldam, 1946

*Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

*Student Solutions Manual for General Chemistry* John McMurry, Robert Fay, 2013-03-07 By Joseph Topich Virginia Commonwealth University This manual for students contains solutions to selected all in chapter problems and even numbered end of chapter problems

**The Hydrogen Atom** Rudolph M. Langer, California Institute of Technology. Division of Physics, Mathematics and Astronomy, 1927

*Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry* Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface

*Chemistry Equations & Answers* Speedy Publishing, 2014-08-26 General chemistry inorganic chemistry organic chemistry and biochemistry are all difficult courses requiring much memorization for the student Essentially there is no easy way to learn formulas and facts This is why chemistry classes are such challenges to students even the best ones However a chemistry equations and answers study guide can help the student When used as a quick reference guide it can be used often to determine the formulas needed for various questions The astute student can cleverly devise ways to make the guide useful for test questions or other circumstances requiring one of the many chemistry equations

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