

CHEMISTRY I EQUATION SHEET

Solubility Rules

Solubility Rules			Form a Compound which is	
Negative Ion Plus Positive Ion				
Any neg. ion + group 1a ions Soluble				
Any neg. ion + Ammonium ion Soluble				
Nitrate + any positive ion Soluble				
Acetate	+	any positive ion except Ag ⁺ or Hg ²⁺	Soluble	
bromide or iodide + Any other positive ion	+	Ag ⁺ , Pb ²⁺ , Hg ²⁺ or Cu ⁺	Not Soluble	
Sulfate	+	Ca ²⁺ , Sr ²⁺ , Ba ²⁺ , Ra ²⁺ , Ag ⁺ or Pb ²⁺	Not Soluble	
	+	Any other positive ion	Soluble	
Sulfide	+	group 1a ions or ammonium	Soluble	
	+	Be ²⁺ , Mg ²⁺ , Ca ²⁺ , Sr ²⁺ , Ba ²⁺ or Ra ²⁺	Soluble	
	+	Any other positive ion	Not Soluble	
Hydroxide	+	group 1a ions or ammonium	Soluble	
	+	Any other positive ion	Not Soluble	
Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion	+	group 1a ions or ammonium	Soluble	
	+	Any other positive ion	Not Soluble	

Equations

D = m/v	c = specific heat
Molarity = $\frac{\text{mol}}{\text{V}}$	d = density
$P_1 V_1 = P_2 V_2$	I = initial amount
$\frac{T_1}{T_2}$	F = final amount
Q = mc ΔT	m = mass
nRTPV =	M = molarity
F = I (1/2) *	n = number of moles & number of half lives
n = Total Time	P = pressure
Half Life Time	R = ideal gas constant
	Q = heat energy
	T = temperature
	V = volume
	X = amount of matter left

Constants

	1 atm = 101.3 kPa Pressure
	1 atm = 760 mm Hg
	R = 8.314 L x kPa / mol°C Universal
Constant or 0.0821 L x atm / mol°C	
Avogadro's number	6.02 X
Specific Heats (J / g°C)	
	Water = 4.180
	Aluminum = 0.895
	Copper = 0.387
	Iron = 0.448

Activity Series in Decreasing Activity

Metals	
Lithium	
Rubidium	
Potassium	
Calcium	
Sodium	
Magnesium	
Aluminum	
Manganese	
Zinc	
Iron	
Nickel	
Tin	
Lead	
Hydrogen	
Copper	
Silver	
Platinum	
Gold	
Nonmetals	
Fluorine	
Chlorine	
Bromine	
Iodine	

Polyatomic Ions

Name	Formula
Ammonium	NH ₄ ⁺
Acetate	C ₂ H ₃ O ₂ ⁻
Carbonate	CO ₃ ⁻²
Bicarbonate or Hydrogen carbonate	HCO ₃ ⁻
Hypochlorite	ClO ⁻
Chlorite	ClO ₂ ⁻
Chlorate	ClO ₃ ⁻
Perchlorate	ClO ₄ ⁻
Hydroxide	OH ⁻
Cyanide	CN ⁻
Nitrate	NO ₃ ⁻
Nitrite	NO ₂ ⁻
Chromate	CrO ₄ ⁻²
Oxalate	C ₂ O ₄ ⁻²
Sulfate	SO ₄ ⁻²
Sulfite	SO ₃ ⁻²
Hydrogen sulfate	HSO ₄ ⁻
Phosphate	PO ₄ ⁻³
Phosphite	PO ₃ ⁻³
Hydrogen phosphate	HPO ₄ ⁻²

Metric Conversions

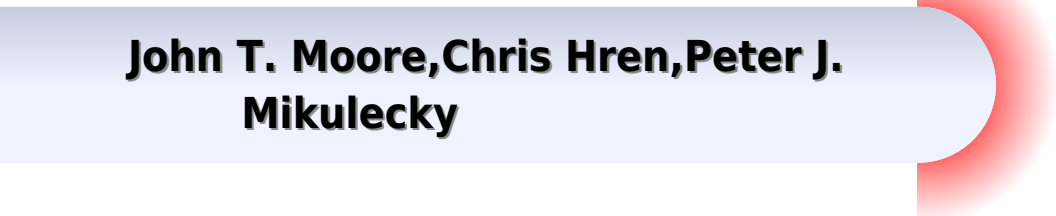
Unit	Abbr.	Meaning
Mega-	M	10 ⁶
kilo-	k	10 ³
hecto-	h	10 ²
deca-	da	10 ¹
Liter	L	1
gram	g	1
meter	m	1
deci-	d	10 ⁻¹
centi-	c	10 ⁻²
milli-	m	10 ⁻³
micro-	μ	10 ⁻⁶
nano-	n	10 ⁻⁹

Calorie Conversions

Water = 1 cal/ g°C
1000 cal = 1 Cal

High School Chemistry Formulas Cheat Sheet

**John T. Moore, Chris Hren, Peter J.
Mikulecky**



High School Chemistry Formulas Cheat Sheet:

Unveiling the Power of Verbal Beauty: An Emotional Sojourn through **High School Chemistry Formulas Cheat Sheet**

In some sort of inundated with displays and the cacophony of fast transmission, the profound energy and psychological resonance of verbal beauty usually diminish in to obscurity, eclipsed by the regular onslaught of noise and distractions. However, nestled within the lyrical pages of **High School Chemistry Formulas Cheat Sheet**, a charming perform of fictional brilliance that pulses with natural thoughts, lies an unforgettable trip waiting to be embarked upon. Published with a virtuoso wordsmith, that mesmerizing opus manuals readers on an emotional odyssey, softly exposing the latent potential and profound impact embedded within the intricate internet of language. Within the heart-wrenching expanse of this evocative examination, we can embark upon an introspective exploration of the book is central themes, dissect their interesting publishing type, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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