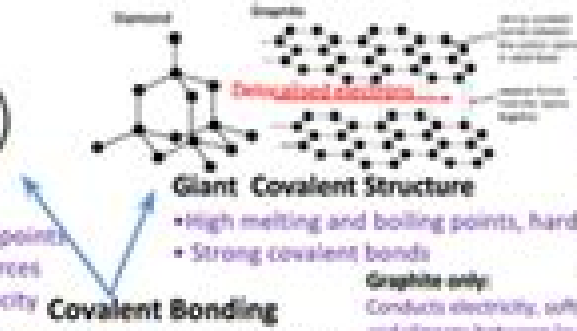
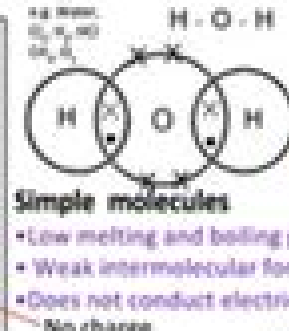




Periodic Table of the Elements

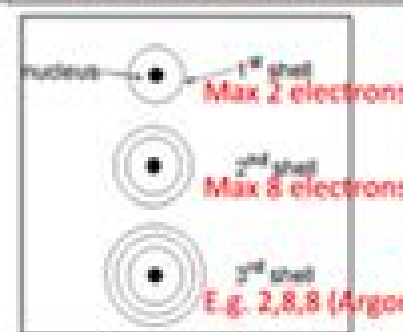
- metals
- non-metals
- metalloids
- transition metals
- alkali metals
- alkaline earth metals
- transition metals
- lanthanoids
- actinoids

An atom:
 Number of Protons = Number of Electrons
 Mass number = Protons + Neutrons

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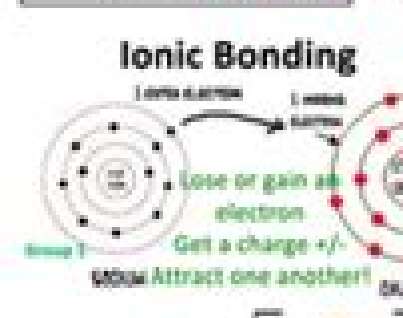
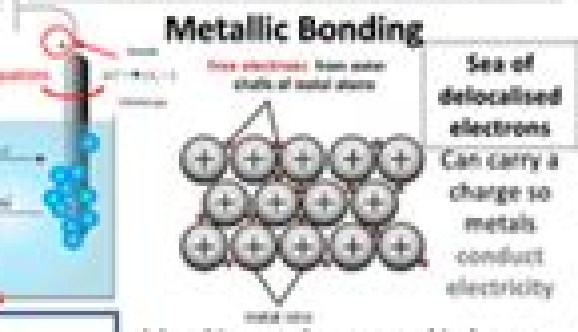
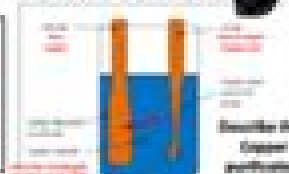
Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
 E.g. Lithium 2,1 is a group 1 element.



OIL RIG

Oxidation is **lose** (of electrons)
 Reduction is **gain** (of electrons)

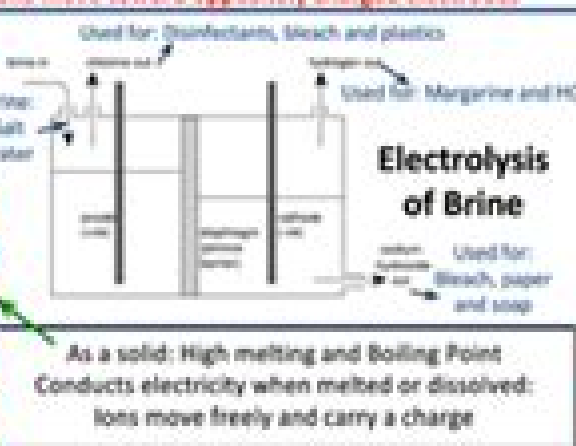
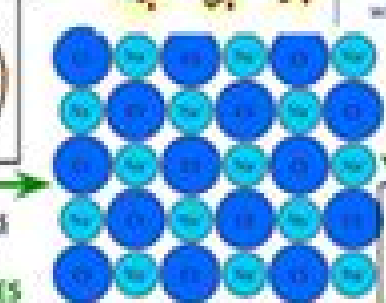


ISOTOPES are atoms of the same element with different numbers of neutrons

When there are 2 possible products The **less reactive** forms at the electrode

Electrolysis needs:

- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply



NANO-TECHNOLOGY
 Means **REALLY REALLY REALLY** small!!!!

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1. Understanding the eBook Gcse Additional Chemistry C2 Revision Booklet
 - The Rise of Digital Reading Gcse Additional Chemistry C2 Revision Booklet
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Additional Chemistry C2 Revision Booklet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Additional Chemistry C2 Revision Booklet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Additional Chemistry C2 Revision Booklet
 - Personalized Recommendations
 - Gcse Additional Chemistry C2 Revision Booklet User Reviews and Ratings
 - Gcse Additional Chemistry C2 Revision Booklet and Bestseller Lists

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 - Gcse Additional Chemistry C2 Revision Booklet eBook Subscription Services
 - Gcse Additional Chemistry C2 Revision Booklet Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Gcse Additional Chemistry C2 Revision Booklet Compatibility with Devices
 - Gcse Additional Chemistry C2 Revision Booklet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gcse Additional Chemistry C2 Revision Booklet
 - Highlighting and Note-Taking Gcse Additional Chemistry C2 Revision Booklet
 - Interactive Elements Gcse Additional Chemistry C2 Revision Booklet
8. Staying Engaged with Gcse Additional Chemistry C2 Revision Booklet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Additional Chemistry C2 Revision Booklet
9. Balancing eBooks and Physical Books Gcse Additional Chemistry C2 Revision Booklet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Additional Chemistry C2 Revision Booklet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gcse Additional Chemistry C2 Revision Booklet
 - Setting Reading Goals Gcse Additional Chemistry C2 Revision Booklet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gcse Additional Chemistry C2 Revision Booklet
 - Fact-Checking eBook Content of Gcse Additional Chemistry C2 Revision Booklet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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