

Can You Find the Side Sides?

In this experiment, boiling acids (H₂SO₄/H₂O) react with hydrocarbons and (HCl) to produce various alkenes, various alcohols, and water. You will determine the moles of reactants used and moles of product produced using both the lab data and stoichiometry and compare the two.

Sample Table:

moles of reactant used and boiling acids	moles
moles of isopropyl alcohol used	moles
moles of isopropyl alcohol	moles
moles of 1-propanol (or isopropanol) produced	moles
moles of isopropyl alcohol used	moles
moles of isopropanol produced	moles

Calculations Questions:

- Calculate the number of moles of boiling acids used in this lab. (10)

$$\text{isopropanol} \times \text{density} \times \text{volume} = \text{mass} \times \frac{1 \text{ mole isopropanol}}{60 \text{ g/mol}}$$
- Calculate the number of moles of product isopropanol produced in this lab. (10)

$$\text{isopropanol} \times \text{density} \times \text{volume}$$
- What is the experimental molar ratio of boiling acids (H₂SO₄/H₂O) to product isopropanol (H₂O)? (10)

$$\frac{\text{isopropanol}_{\text{moles}}}{\text{isopropanol}_{\text{moles}}}$$
- What is balanced equation for the reaction that took place in this experiment? (10)

$$\text{isopropanol} + \text{H}_2\text{SO}_4 \rightarrow \text{isopropanol} + \text{H}_2\text{O}$$
- According to the balanced equation, what is the theoretical molar ratio of boiling acids to product isopropanol? (10)

$$1:1$$
- What the experimental molar ratio used by this reaction for theoretical molar ratio? Or not, give some possible reasons. (10)

$$\frac{\text{isopropanol}_{\text{moles}}}{\text{isopropanol}_{\text{moles}}} \text{ used as 1:1 ratio, but experimental isopropanol}_{\text{moles}}$$

Episode 801 Chemistry Answers

Pierre Vogel, Kendall N. Houk



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

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