

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kJoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kJoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	<p>Across a period (→) atomic radius tends to _____.</p> <p>Across a period (→) first ionization energy tends to _____.</p> <p>Across a period (→) electronegativity tends to _____.</p>						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	<p>Down a Group (↓) atomic radius tends to _____.</p> <p>Down a Group (↓) first ionization energy tends to _____.</p> <p>Down a Group (↓) electronegativity tends to _____.</p>						
Rb							
Cs							

Laboratory Exercise On Graphing Periodic Trends

Answers

Terry C. Jones



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1. Understanding the eBook Laboratory Exercise On Graphing Periodic Trends Answers
 - The Rise of Digital Reading Laboratory Exercise On Graphing Periodic Trends Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Laboratory Exercise On Graphing Periodic Trends Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Laboratory Exercise On Graphing Periodic Trends Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Laboratory Exercise On Graphing Periodic Trends Answers
 - Personalized Recommendations
 - Laboratory Exercise On Graphing Periodic Trends Answers User Reviews and Ratings
 - Laboratory Exercise On Graphing Periodic Trends Answers and Bestseller Lists
5. Accessing Laboratory Exercise On Graphing Periodic Trends Answers Free and Paid eBooks
 - Laboratory Exercise On Graphing Periodic Trends Answers Public Domain eBooks
 - Laboratory Exercise On Graphing Periodic Trends Answers eBook Subscription Services
 - Laboratory Exercise On Graphing Periodic Trends Answers Budget-Friendly Options
6. Navigating Laboratory Exercise On Graphing Periodic Trends Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Laboratory Exercise On Graphing Periodic Trends Answers Compatibility with Devices
 - Laboratory Exercise On Graphing Periodic Trends Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Laboratory Exercise On Graphing Periodic Trends Answers
 - Highlighting and Note-Taking Laboratory Exercise On Graphing Periodic Trends Answers
 - Interactive Elements Laboratory Exercise On Graphing Periodic Trends Answers
8. Staying Engaged with Laboratory Exercise On Graphing Periodic Trends Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Laboratory Exercise On Graphing Periodic Trends Answers
9. Balancing eBooks and Physical Books Laboratory Exercise On Graphing Periodic Trends Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Laboratory Exercise On Graphing Periodic Trends Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Laboratory Exercise On Graphing Periodic Trends Answers
 - Setting Reading Goals Laboratory Exercise On Graphing Periodic Trends Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Laboratory Exercise On Graphing Periodic Trends Answers
 - Fact-Checking eBook Content of Laboratory Exercise On Graphing Periodic Trends Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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