

# Electron Configuration of Main Group Cations

Electron configuration

**Na**  $1s^2 2s^2 2p^6 3s^1$

Orbital diagram

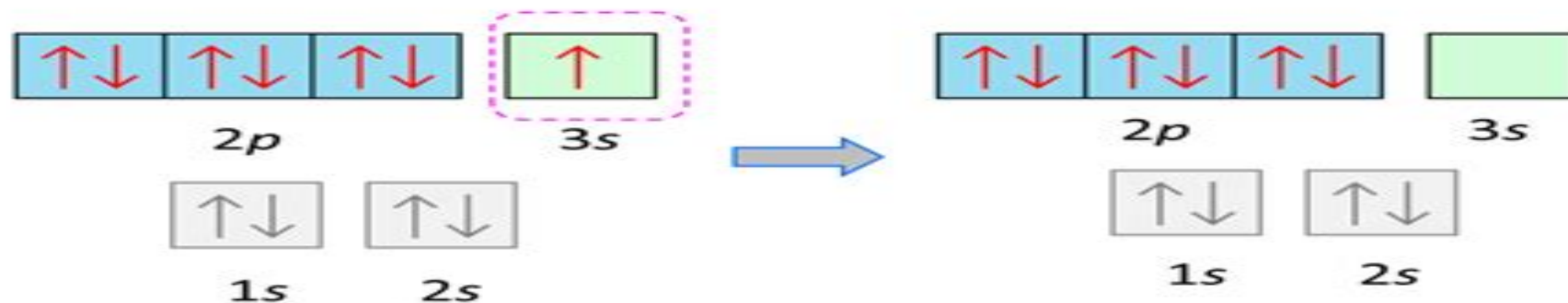


Focus on the valence shell

The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital –  $3s^1$



**Na**  $1s^2 2s^2 2p^6 3s^1$

**Na<sup>+</sup>**  $1s^2 2s^2 2p^6$

↓ Same as

**Ne**  $1s^2 2s^2 2p^6$



ChemistrySteps.com

Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

# Lab Electron Configurations Of Atoms And Ions

**Ying Liu**



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**Basic Chemistry** Steven S. Zumdahl, 1999 For a full description see catalog entry for Zumdahl Introductory Chemistry A Foundation 4 e *Technical Abstract Bulletin* Defense Documentation Center (U.S.), 1963 **Government Reports**

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## Unveiling the Magic of Words: A Overview of "**Lab Electron Configurations Of Atoms And Ions**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Lab Electron Configurations Of Atoms And Ions**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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