

23. When red litmus paper is added to lime water
 (a) it changes to blue
 (b) it remains red
 (c) it becomes green
 (d) it changes into yellow colour.
24. Solution A turns blue litmus red, evolves H_2 gas with Zn metal and gives out CO_2 on treatment with Na_2CO_3 . Solution B turns red litmus blue, liberates H_2 with Zn metal and does not react with Na_2CO_3 . Which of the following option is correct?
 (a) Solution A is NaOH; solution B is HCl.
 (b) Solution A is HCl; solution B is NaOH.
 (c) Solution A is NaOH; solution B is Na_2CO_3 .
 (d) Solution A is Na_2CO_3 ; solution B is NaOH.
25. A dilute solution of sodium carbonate was added to two test tubes containing dilute HCl (A) and dilute NaOH (B) respectively. The correct observation was
 (a) a brown coloured gas liberated in test tube A
 (b) a brown coloured gas liberated in test tube B
 (c) a colourless gas liberated in test tube A
 (d) a colourless gas liberated in test tube B.

26. Alkalies are
 (a) acids dissolved in water
 (b) bases dissolved in water
 (c) a mixture of acid and base
 (d) a diluted solution of HCl.
27. The correct formula for calculation of pH is
 (a) $+\log [H^+]$ (b) $-\log [OH^-]$
 (c) $-\log [H^+]$ (d) $-\log [H^+] + \log [OH^-]$
28. Antacids are taken for _____ reaction, when excess acid is formed in stomach.
 (a) combination (b) decomposition
 (c) photochemical (d) neutralisation
29. An indicator shows red colour in acid solution and yellow colour in basic solution. The indicator is
 (a) phenolphthalein (b) methyl orange
 (c) litmus (d) China rose.
30. The number of ionisable hydrogen ions in H_2SO_4 and HCl are respectively
 (a) 2, 3 (b) 3, 1
 (c) 2, 1 (d) 1, 1

Unboxing Apply Your Concept

1. (b) : Sodium carbonate is a basic salt. When dissolved in water, it gives an alkaline (or basic) solution.
2. (d) : The given solution is basic in nature. When excess of HCl is added, it becomes acidic.
3. (a) : The sample Y with $pH = 13$ is a strongly basic compound. Hence, it is NaOH.
4. (b) : The increasing order of pH is
 Highly acidic \rightarrow Moderately acidic \rightarrow Neutral \rightarrow
 Moderately basic \rightarrow Highly basic (alkaline)
 For the given substances, the increasing order of pH will be HCl, CH_3COOH , NaOH.
5. (b) : Soft drinks are moderately acidic. In this range of pH, the pH paper will show yellow-orange colour.
6. (b) : In acidic solution, blue litmus paper changes to red and in basic solution, red litmus paper changes to blue. Hence, blue litmus paper first changes its colour to red and then to blue.
7. (a) : $[H^+] = 1.0$, $pH = -\log [H^+] = -\log (1)$
 So, $pH = 0$
8. (c) : The concept of pH is not applicable to dry solid substances. Therefore, pH of solid sodium bicarbonate cannot be found directly using a pH paper.

9. (b) : According to the experiment performed,

Solutions	I	II	III	IV
Colour of pH paper :	Red	Blue	Light green	Orange
Conclusion :	Highly acidic	Highly alkaline	Neutral	Moderately acidic

So, I and IV are acidic.

10. (b) : A : Blue Highly alkaline solution Sodium carbonate solution
 B : Orange Moderately acidic solution Acetic acid
11. (d) : Universal indicators are mixture of many indicators.
12. (d) : The correct method to find pH of a solution is to put the solution on a pH paper using a dropper.
13. (d) : Light green (A) : Almost neutral
 Dark red (B) : Highly acidic
 Light orange (C) : Slightly acidic
 Dark blue (D) : Strong basic
 Here order of pH: $D > A > C > B$
14. (a) : Zinc reacts with dil. H_2SO_4 to produce hydrogen gas.
 $Zn_{(s)} + H_2SO_{(aq)} \rightarrow ZnSO_{(aq)} + H_{2(g)} \uparrow$

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