
Identify the order of the change that best completes the statement or answers the question.

- _____ 7. In many compounds, atoms of main-group elements form ions so that the number of electrons in their outermost energy levels of each ion is:
- a. 2. c. 8.
b. 6. d. 10.
- _____ 8. An atom
- a. is an ion with a negative charge.
b. attracts ions with negative charges.
c. results when an alkaline-earth metal loses one of its two valence electrons.
d. has more protons than electrons.
- _____ 9. When the octet rule is satisfied, the outermost _____ are filled.
- a. all orbitals c. s and p orbitals
b. s and p orbitals d. all s and p orbitals
- _____ 10. The elements of the _____ group satisfy the octet rule without forming compounds.
- a. halide b. alkali metal
c. noble gas d. alkaline earth metal
- _____ 11. Which gas atom has full s and p orbitals in its outermost energy level?
- a. It is highly reactive only with alkali metals.
b. It is highly reactive only with halogens.
c. It can be combined with most elements.
d. It has a stable octet and is unreactive.
- _____ 12. In a crystal of an ionic compound, each cation is surrounded by
- a. molecules. b. dipoles.
c. positive ions. d. anions.
- _____ 13. The lithium(I) ion and sodium(I) ion
- a. have the same charge.
b. are polyatomic ions.

Ionic And Metallic Bonding Chapter 7 Worksheet Answers

R Pring



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