

Beantwortet

1	<p>Welche der folgenden Aussagen sind richtig (R) oder falsch (F)?</p> <p>a) ΔH ist die Enthalpie, ΔS die Entropie, ΔG die freie Enthalpie. b) ΔH ist die Enthalpie, ΔS die Entropie, ΔG die freie Enthalpie. c) ΔH ist die Enthalpie, ΔS die Entropie, ΔG die freie Enthalpie. d) ΔH ist die Enthalpie, ΔS die Entropie, ΔG die freie Enthalpie. e) ΔH ist die Enthalpie, ΔS die Entropie, ΔG die freie Enthalpie. f) ΔH ist die Enthalpie, ΔS die Entropie, ΔG die freie Enthalpie.</p>
2	<p>Die freie Enthalpie ΔG ist die Differenz aus Enthalpie ΔH und Entropie ΔS. a) $\Delta G = \Delta H - T \Delta S$ b) $\Delta G = \Delta H + T \Delta S$</p>
3	<p>Die freie Enthalpie ΔG ist die Differenz aus Enthalpie ΔH und Entropie ΔS. a) $\Delta G = \Delta H - T \Delta S$ b) $\Delta G = \Delta H + T \Delta S$</p>
4	<p>Die freie Enthalpie ΔG ist die Differenz aus Enthalpie ΔH und Entropie ΔS. a) $\Delta G = \Delta H - T \Delta S$ b) $\Delta G = \Delta H + T \Delta S$</p>
5	<p>Die freie Enthalpie ΔG ist die Differenz aus Enthalpie ΔH und Entropie ΔS. a) $\Delta G = \Delta H - T \Delta S$ b) $\Delta G = \Delta H + T \Delta S$</p>

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