

# Unit 6- Mole Conversions (NB p. 49-51)

Vocabulary: Avogadro's number, molar mass, mole ratio, conversion factor, dimensional analysis

Equations & Constants: Avo#:  $6.02 \times 10^{23}$  particles = 1 mole

Practice Problems:

- C 1. The mass of 1 mol of chromium is  
a. 12 g. ☐ b. 198 g. ☐ c. 51.996 g. ☐ d.  $6.02 \times 10^{23}$  g.

$$\frac{1 \text{ mole Cr}}{1 \text{ mole Cr}} \times \frac{g \text{ Cr}}{1 \text{ mole Cr}}$$

- D 2. The mass of a sample containing 3.5 mol of silicon atoms is  
a. 28 g. ☐ b. 35 g. ☐ c. 72 g. ☐ d. 98 g. ☐

$$\frac{3.5 \text{ mol Si}}{1 \text{ mol}} \times \frac{g \text{ Si}}{1 \text{ mol}}$$

- A 3. A prospector finds 39.39 g of gold. She has  
☐ a.  $1.20 \times 10^{23}$  atoms. ☐ c.  $4.30 \times 10^{23}$  atoms.  
☐ b.  $2.30 \times 10^{23}$  atoms. ☐ d.  $6.02 \times 10^{23}$  atoms.

- B 4. A sample of tin contains  $3.01 \times 10^{23}$  atoms. The mass of the sample is  
☐ a. 3.01 g. ☐ c. 72.6 g. ☐ b. 59.3 g. ☐ d. 11 g.

$$\frac{3.01 \times 10^{23} \text{ atoms}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mol}}{1 \text{ mol}} \times \frac{118.71 \text{ g}}{1 \text{ mol}}$$

- B 5. The mass of a sample of nickel is 176.07 g. It contains  
☐ a.  $1.7607 \times 10^{24}$  atoms. ☐ c.  $5.869 \times 10^{23}$  atoms.  
☐ b.  $1.806 \times 10^{24}$  atoms. ☐ d.  $5.869 \times 10^{24}$  atoms.

$$\frac{176.07 \text{ g Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol}}{1 \text{ mol}} \times \frac{6.02 \times 10^{23}}{1 \text{ mol}}$$

- B 6. What is the mass of a 6.267-mol sample of sodium hydroxide?  
☐ a. 40.00 g ☐ c. 150.4 g ☐ b. 250.7 g ☐ d. 6.382 g

$$\text{NaOH} \\ (6.267) \times (40 \text{ g/mol})$$

7. How many atoms of hydrogen are present in 3.18 g of ammonia?  
☐ a.  $1.12 \times 10^{23}$  ☐ c.  $5.74 \times 10^{24}$  ☐ b.  $6.33 \times 10^{23}$  ☐ d.  $3.37 \times 10^{23}$

$$\frac{3.18 \text{ g}}{17 \text{ g/mol}}$$

8. How many moles are in 2.3 g of a sample of potassium chlorate?

9. How many carbon atoms are in 9.22 moles of butane,  $\text{C}_4\text{H}_{10}$ ?

Convert grams X to moles X

$$\frac{\text{grams X}}{1 \text{ mole}} \times \frac{1 \text{ mole}}{\text{grams (MM) X}}$$

Convert atoms of X to moles of X

$$\frac{\text{atoms X}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole X}}{1 \text{ mole X}}$$

Convert atoms of X to grams of X

Convert mole X to grams X

$$\frac{\text{mole X}}{1 \text{ mole X}} \times \frac{\text{grams (MM) X}}{1 \text{ mole X}}$$

Convert moles X to atoms of X

$$\frac{\text{moles X}}{1 \text{ mole X}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole X}}$$

Convert grams of X to atoms of X

How many atoms of X in grams of XY?

How many atoms of X in moles of XY?

# Grayslake Chemistry Review Packet Answer Key

**M Woodhall**



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