

Gravimetric Analysis of a Chloride Salt (G)

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Lab section: LW3

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Purpose:

The purpose of the lab is to find the amount of chloride present in any unknown soluble salt, by showing and explaining the techniques present in gravimetric analysis.

Theory:

- i. This is the equation that describes the reaction between silver ion and chloride ion to form silver chloride:



- ii. AgCl solubility in water and the K_{sp} :

"Silver chloride is very insoluble, so we can assume that the above reaction goes to completion."¹

$$K_{sp} = 1.6 \times 10^{-10}$$

- iii. AgCl is precipitated in an acidic medium to reduce the chance of acid ions interfering with the results. These ions will form co-precipitates with silver in acid. Due to these co-precipitates, the results would be higher.
- iv. In order to ensure the precipitate does not go through the filter paper, which would give us lower results. To prevent this from happening, it is heated to make coagulate it into a clumpy mass. Heating the precipitate makes it coagulate, making it much harder for any precipitate to go through the filter paper.
- v. Silver chloride decomposes into silver and chloride in the presence of light, which is described by the following equation:



If this photodecomposition takes place in air, it will result in lower results since chlorine will be released as a gas. If it takes place in the presence of excess silver ion in aqueous solution, it will result in higher results, due to chlorine reacting with the silver to form silver chloride and solid silver from the decomposition remaining in the precipitate, which will make the results high. The following chemical equation is a result of chlorine reacting with silver to form silver chloride:

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