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#### CHEMISTRY PAPER 2 (233/2)

#### FORM FOUR (4)

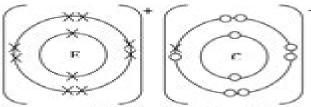
#### TIME: 2 HOURS

#### MARKING SCHEME

 $(1 \, \text{mk})$ (1, mik)

- 1 (a) Alkaline-earth metals
  - (b) B
  - (c) (i) A<sub>2</sub>B<sub>3</sub> (½mk)
- (ii) H<sub>2</sub>SO<sub>4</sub>
  - (3/2mk).





Distribution of electrons (1mk) and charges (1mk)

No label (0mk)

(e) D has a smaller atomic size than C//C has larger atomic size than D (1mk)

D has more protons than C//D has stronger nuclear charge than C (1mk)

(f) H (½mk). Loses electrons most readily/largest atomic radius/weakest nuclear charge (explanation ½mk)

(g) 2A(s) + 6HCl(aq) → 2ACl3(aq) + 3H2(g) 1mk

Moles of A = 
$$\frac{0.081}{27}$$
 = 0.003moles ½mk

Mole ratio 1:3

Moles of HCl = 0.003x3 = 0.009 moles  $\frac{1}{2}$ mk

0.009×1000 ½mk = 0.45M ½mk

(h) The melting point increase with increase in atomic number. 1mk

This is due to increase in the strength of van der waal's forces (inter-molecular forces) 1mk

# **Marking Scheme For Chemistry Paper2 2002**

F Rizvi

# Marking Scheme For Chemistry Paper 22002:

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# **Table of Contents Marking Scheme For Chemistry Paper 22002**

- 1. Understanding the eBook Marking Scheme For Chemistry Paper 22002
  - The Rise of Digital Reading Marking Scheme For Chemistry Paper 22002
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Marking Scheme For Chemistry Paper 22002
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Marking Scheme For Chemistry Paper2 2002
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Marking Scheme For Chemistry Paper 22002
  - Personalized Recommendations
  - Marking Scheme For Chemistry Paper2 2002 User Reviews and Ratings
  - Marking Scheme For Chemistry Paper2 2002 and Bestseller Lists
- 5. Accessing Marking Scheme For Chemistry Paper 22002 Free and Paid eBooks
  - Marking Scheme For Chemistry Paper2 2002 Public Domain eBooks
  - Marking Scheme For Chemistry Paper2 2002 eBook Subscription Services
  - Marking Scheme For Chemistry Paper2 2002 Budget-Friendly Options
- 6. Navigating Marking Scheme For Chemistry Paper2 2002 eBook Formats

- o ePub, PDF, MOBI, and More
- Marking Scheme For Chemistry Paper 22002 Compatibility with Devices
- Marking Scheme For Chemistry Paper2 2002 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Marking Scheme For Chemistry Paper 22002
  - Highlighting and Note-Taking Marking Scheme For Chemistry Paper 22002
  - Interactive Elements Marking Scheme For Chemistry Paper 22002
- 8. Staying Engaged with Marking Scheme For Chemistry Paper 22002
  - Joining Online Reading Communities
  - o Participating in Virtual Book Clubs
  - Following Authors and Publishers Marking Scheme For Chemistry Paper 22002
- 9. Balancing eBooks and Physical Books Marking Scheme For Chemistry Paper 22002
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Marking Scheme For Chemistry Paper 22002
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Marking Scheme For Chemistry Paper 22002
  - Setting Reading Goals Marking Scheme For Chemistry Paper 22002
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Marking Scheme For Chemistry Paper 22002
  - Fact-Checking eBook Content of Marking Scheme For Chemistry Paper 22002
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
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