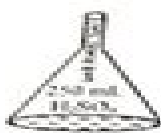


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটীক্ষার কপূর, এখানে সেরে পটীক্ষা, আরও কলেজের শিক্ষার্থীর কলেজসমূহের পটীক্ষা এবং মতামত প্রদানের ক্ষমতা
পূর্ণাঙ্গ উত্তর অধ্যয়নিক সেবা প্রদান। এগুলোর অনুশীলন করলে তুমি এ অধ্যয়ন থেকে সেরে পটীক্ষা পূর্ণাঙ্গ প্রদান।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেলস পানিদ্রুত লবণের প্রকণতা, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
- A-পাত্রের বিন্দুমান এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা প্রকণ হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণ কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলের প্রকণ বলে।

কেলস পানিদ্রুত লবণের প্রকণতা তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে প্রকণতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কেলস পানিদ্রুত লবণ নিষ্কৃতি হয়। তখন প্রকণতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে প্রকণিত করার সময় বিশ্লেষণ প্রক্রিয়াটি তাপমাত্রা বৃদ্ধি 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিশ্লেষণ প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে প্রকণতার হ্রাস ঘটে।

২

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
অতএব, প্রকণিত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকণিত H_2SO_4 এর ভর, $w = ?$

প্রকণ-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রকণ-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} \times 1 = 0.077 \text{ mol}$
 $= 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকণের সাথে পাত্র-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ প্রকণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর প্রকণ প্রকণ করে।

সি. কে. ১০/১০/১০

- নোড কাকে বলে?
- Cl_2 অপোলার— ব্যাখ্যা করো।
- উদাহরণের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদাহরণের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা এক হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিন্দু ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিন্দু সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌলিক অণুর ক্ষেত্রে এহুণ ঘাটক।

উদাহরণে বর্ণিত এলিটটি হলো 2% $\left(\frac{W}{V}\right) \text{ HCl}$

2% $\left(\frac{W}{V}\right) \text{ HCl}$ এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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Donna Kerr, Sandra Wilson



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