

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol, C_2H_5OH , are present in 65 mL of a 1.5 M solution?
2. How many liters of a 6.0 M solution of acetic acid CH_3COOH , contain 0.0030 moles of acetic acid?
3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?
4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?
5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)
6. Convert the following Molarities to Normalities.
 - a. 2.5 M HCl = _____ N
 - b. 1.4 M H_2SO_4 = _____ N
 - c. 1.0 M NaOH = _____ N
 - d. 0.5 M $Ca(OH)_2$ = _____ N
7. A commonly purchased disinfectant is a 3.0% (by mass) solution of hydrogen peroxide (H_2O_2) in water. Assuming the density of the solution is 1.0 g/cm³, calculate the molarity and molality of H_2O_2 .
8. A solution is made by dissolving 25 g of NaCl in enough water to make 1.0 L of solution. Assume the density of the solution is 1.0 g/cm³. Calculate the molarity and molality of the solution.
9. What is the mass percent of a solution that contains 152 g of KNO_3 in 7.86 L of water?
10. A solution has been prepared to be 9.00 % (by mass) glucose. For every 100.0 grams of solution present there are _____ grams of glucose.
11. How many grams of KBr are contained in 250. grams of a 6.25 % KBr solution?
12. What is the concentration of each type of ion and total concentration of ions in a 0.375 M Ammonium phosphate solution?
13. How many moles of chloride ions are in 1.50 L of 4.15 M zinc chloride?

ANSWERS

1. 0.098 moles of alcohol
2. 5.0×10^{-4} L solution
3. 5.9×10^{-2} g solution
4. 2.8M
5. 6.6 M
6.
 - a. 2.5 N
 - b. 2.8 N
 - c. 1.0 N
 - d. 1 N
7. 0.88M, 0.91m
8. 0.43M, 0.43m
9. 1.90 % by mass of KNO_3
10. 9.00
11. 15.6 g KBr
12.
 - a. 1.13 M of $(NH_4)^{+1}$ and 0.375 M of $(PO_4)^{-3}$
 - b. 1.51 M of ions
13. 12.45 moles of Cl^{-1}

Holt Chemistry Section Concentration And Molarity Answers

J Elliott



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Holt Chemistry Holt Rinehart & Winston,2003-01-24 **Principles of Modern Chemistry** David W. Oxtoby,H. Pat Gillis,Laurie J. Butler,2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses **PRINCIPLES OF MODERN CHEMISTRY** continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Modern Chemistry**,1990 *Ebook: Chemistry* Julia Burdge,2014-10-16 Chemistry Third Edition by Julia Burdge offers a clear writing style written with the students in mind Julia uses her background of teaching hundreds of general chemistry students per year and creates content to offer more detailed explanation on areas where she knows they have problems With outstanding art a consistent problem solving approach interesting applications woven throughout the chapters and a wide range of end of chapter problems this is a great third edition text Holt Chemistry File ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket **Holt Chemistry** Ralph Thomas Myers,2004 Modern Chemistry H. Clark Metcalfe,Joseph F. Castka,1966

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