

Aim:

To find pH of the following samples by using paper/universal indicator

- (a) Dilute hydrochloric acid
- (b) Dilute NaOH solution
- (c) Dilute ethanoic acid solution
- (d) Lemon juice
- (e) Water
- (f) Dilute sodium bicarbonate solution

Theory:

• pH is the measure of the hydrogen ion concentration $[H^+]$ of a solution

$$pH = -\log_{10} [H^+]$$

• Acids release H^+ ions when dissolved in water

• Bases release OH^- ions when dissolved in water

If $pH < 7$ it is a acidic solution

If $pH > 7$ it is a basic solution

If $pH = 7$ then it is neutral

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