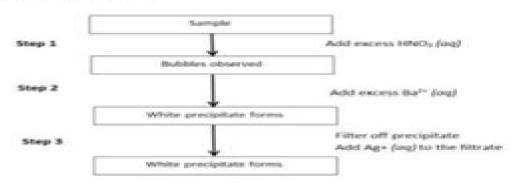




Chemistry Form 10

(6 points) The flow diagram shows tests that can be used to identify carbonate, chloride and sulfate ions
present in a sample.



Write balanced identification equations for all reactions in these steps in ionic and overall forms.

(11 points) Bone consists of a complex mixture of calcium salts, such as calcium carbonate, and other
material. If bone is strongly heated, the calcium carbonates turn into calcium oxide, CaO. This can be
separated from the other material by dissolving it in dilute hydrochloric acid, to form calcium chloride,
CaCl₂, solution.

a. Write complete and balanced chemical equations showing state symbols for these reactions. 2 points

b. An experiment showed that 14.4 g of calcium oxide were made from a sample of bone of mass 50.0 g. Calculate the % of calcium by mass in the bone. Calculate the not used amount of hydrochloric acid in the reaction in g, if obtained calcium oxide was completely dissolved in 100 ml 20 % of dilute hydrochloric acid. Hydrochloric acid solution density is 1.1 g/ml. 6 points

c. Strontium compounds are incorporated into bones because their chemistry is similar to that of calcium compounds. For example, both strontium and calcium form ions with the same charge. Explain, in terms of the Periodic Table, the electronic structures of the atoms why calcium and strontium (and their compounds) have similar chemical properties. 3 points

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