Give the oxidation numbers of all the elements in the following molecules and ions:

a. 50, 502, 503, 5032, 5042.

b. clo2, clo3, clo2, clo3, clo4

c. N2O, NO, NO2, N2O4, N2O5, NO2, NO3

2. Determine the oxidation number of the sulfur atom:

a. H2S b. S c. H2SO4 d. S2. e. HS. f. SO2 q. SO3

Indicate the oxidation number of phosphorus in each of the following compounds:

a. HPO₃ _____

d. H₃PO₄

b. H₃PO₂ ____

e. H₄P₂O₇ _____

c. H₃PO₃ _____

f. H₅P₃O₁₀ _____

4. Give oxidation numbers for the underlined atoms in these molecules and ions:

a. Cs2O

f. CIF₃

k. MoO42-

b. PtCl62-

q. H₃AsO₃

I. MnO4"

c. Calz

h. 5bF6

m. PtCl42-

d. SnF2

i. <u>Ti</u>O₂

n. <u>O</u>2

e. <u>Al</u>₂O₃

J. P4

Q₃

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