

physique

quantité de matière :

Le nombre d'entités par mole est appelée constante d'avogadro :

$$N_A = 6,02 \times 10^{23} \text{ mol}^{-1}$$

n = nombre de paquets contenant N_A

$$\text{donc } N = n \times N_A$$

↳ en moles (mol)

Pour un échantillon :

$$M = \frac{m}{n} ; \text{ masse échantillon : } m = n \times M$$

M = masse molaire ($\text{g} \cdot \text{mol}^{-1}$)

m = masse (g)

n = quantité matière (mol)

masse molaire :

- pour un atome : la masse molaire est donnée dans le tableau périodique
- pour une molécule : c'est la somme des masses molaires atomiques de tous les atomes composant la molécule
- pour un ion : elle est égale à la masse molaire atomique de l'élément correspondant

Rappel :

hydrogène : $1 \text{ g} \cdot \text{mol}^{-1}$

carbone : $12 \text{ g} \cdot \text{mol}^{-1}$

azote : $14 \text{ g} \cdot \text{mol}^{-1}$

oxygène : $16 \text{ g} \cdot \text{mol}^{-1}$

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Wenbin Ji



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