

# DAT General Chemistry Equation Sheet

## General Concepts

Density:  $(d = m/v)$

Dilutions:  $M_1V_1 = M_2V_2$  (or  $C_1V_1 = C_2V_2$ )

% Error Calculations

$$\% \text{ Error} = \frac{\text{Actual} - \text{theoretical}}{\text{theoretical}} \times 100$$

Spectrophotometer:  $Abs = \epsilon cl$

$Abs =$  Absorbance

$\epsilon =$  molar extinction coefficient (molar absorptivity)

$l =$  path length

## Atomic Structure

Absorption/Emission Line Spectra:  $\Delta E = E_{\text{photon}}$

Energy of a photon:  $E_{\text{photon}} = hf = \frac{hc}{\lambda}$

$$c = 3.0 \times 10^8 \text{ m/s}$$

$$h = 6.63 \times 10^{-34} \text{ J-s}$$

Photoelectric Effect:  $KE_{\text{electron}} = E_{\text{photon}} - \phi$

$\phi =$  work function (minimum energy to ionize electron)

## Chemical Kinetics and Equilibrium

General Rate Law:  $\text{rate} = k[A]^m[B]^n$

Arrhenius Equation:  $k = Ae^{-E_a/RT}$   $E_a =$  activation energy

Equilibrium Constant:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

$R = 0.0821 \text{ L atm/mol K}$

## Solutions

### Concentration

$$\text{Molarity (M)} = \frac{\text{moles}_{\text{solute}}}{V_{\text{solution}}}$$

$$\text{molality (m)} = \frac{\text{moles}_{\text{solute}}}{kg_{\text{solvent}}}$$

Henry's Law:  $P_A = k_H[A]$

### Colligative Properties

Freezing Point Depression:  $\Delta T_F = -iK_F m$

Boiling Point Elevation:  $\Delta T_B = iK_B m$

Raoult's Law (\* = pure):  $P_A = X_A P_A^*$

Osmotic Pressure:  $\pi = iMRT$

$K_F/K_B = i \cdot P/B.P$   
depression/elevation constant  
 $X_A =$  mole fraction of A  
 $i =$  van't Hoff Factor  
 $R = 0.0821 \text{ L atm/mol K}$   
 $M =$  molarity,  $m =$  molality

## Gases/Gas Laws

### Average Kinetic Energy

$KE_{\text{avg}} = \frac{3}{2} RT$  (per mole of gas)

Root-mean-square speed:  $v = \sqrt{\frac{3RT}{M_m}}$

### Gas Laws

Ideal Gas Law:  $PV = nRT$

Boyle's Law:  $P \propto \frac{1}{V}$

Charles' Law:  $V \propto T$

Avogadro's Law:  $V \propto n$

Combined Gas Law:  $P_1V_1 = \frac{P_2V_2}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$

### Dalton's Law of Partial Pressures:

$P_{\text{Tot}} = P_A + P_B + \dots$

$P_A = X_A \cdot P_{\text{Tot}}$

$X_A =$  mol fraction A

### Grahms Law of Effusion:

$$\frac{r_1}{r_2} = \sqrt{\frac{M_{m2}}{M_{m1}}}$$

$R = 0.0821 \text{ L atm/mol K}$   
 $n =$  moles  
 $M_m =$  molar mass  
 $r =$  rate of effusion  
 $a$  and  $b$  are constants specific to a given gas

## Acid-Base

### pH Scale

$pH = -\log [H^+]$   $[H^+] = 10^{-pH}$

$pOH = -\log [OH^-]$   $[OH^-] = 10^{-pOH}$

$pH + pOH = 14$   $[H^+][OH^-] = 1 \cdot 10^{-14}$

### Weak Acids:

$$K_a = \frac{[H_3O^+][A^-]}{[HA]} [H^+] = \sqrt{K_a [HA]}$$

### Weak Bases:

$$K_b = \frac{[OH^-][MA]}{[A^-]} [OH^-] = \sqrt{K_b [B]}$$

$$K_a + K_b = K_w$$

Buffer:  $pH = pK_a + \log \frac{[A^-]}{[HA]}$

Titration/Neutralization:  $n_a M_a V_a = n_b M_b V_b$

# Formula Sheet For Gen Chem 1 Chm 2045

**Helmut J. Fischbeck, Kurt H. Fischbeck**

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