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Possibly useful info:  $R=0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K}$ ;  $R=8.314 \text{ J/mol}\cdot\text{K}$ ;  $1 \text{ J} = 1 \text{ kg}\cdot\text{m}^2/\text{s}^2$ ;  $h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$ ;  
 $c = 2.998 \times 10^8 \text{ m/s}$ ; specific heat capacity ( $\text{H}_2\text{O}$ ) =  $4.184 \text{ J/g}\cdot^\circ\text{C}$

- Calculate the concentration (in molarity) of a NaOH solution if 25.0 mL of the solution are needed to neutralize 17.4 mL of a 0.312 M HCl solution.  
 (1) 0.146 M (2) 0.217 M (3) 0.446 M (4) 0.517 M (5) 0.733 M
- A solid will float on any liquid that is more dense than the solid. The volume of a piece of calcite weighing 35.6 g is  $12.9 \text{ cm}^3$ . On how many of the following liquids will calcite float? Carbon tetrachloride ( $d=1.60 \text{ g/cm}^3$ ); methylene bromide ( $d=2.50 \text{ g/cm}^3$ ); tetrabromoethane ( $d=2.96 \text{ g/cm}^3$ ); methylene iodide ( $d=3.33 \text{ g/cm}^3$ )  
 (1) 1 (2) 2 (3) 3 (4) 4 (5) 0
- What is the molecular formula of a gaseous compound (composed solely of C, H, and Cl atoms) whose mass is 0.568 g at  $26^\circ\text{C}$  and 740 torr in a 120 mL flask?  
 (1)  $\text{NH}_4\text{NO}_3$  (2)  $\text{C}_2\text{H}_5\text{Cl}$  (3)  $\text{C}_2\text{H}_4\text{Cl}_2$  (4)  $\text{CHCl}_3$  (5)  $\text{CCl}_4$
- For the  $\text{XeF}_4$  molecule, what is the hybridization on Xe, the number of lone pairs (LP), and does it possess a dipole moment (DM) or not?  
 (1)  $d^2sp^3$  (or  $sp^3d^2$ ), two LP, DM = 0 (2)  $sp^3$ , zero LP, DM = 0  
 (3)  $dsp^3$  (or  $sp^3d$ ), one LP, DM  $\neq 0$  (4)  $dsp^3$  (or  $sp^3d$ ), two LP, DM  $\neq 0$   
 (5)  $d^2sp^3$  (or  $sp^3d^2$ ), two LP, DM  $\neq 0$
- Which of the following would be predicted to have the smallest size and why?  
 (1)  $\text{Mg}^{2+}$  because it has the most metallic character  
 (2)  $\text{Al}^{3+}$  because it has the highest effective nuclear charge  $Z_{\text{eff}}$   
 (3) Ne because it is inert and has a filled shell  
 (4)  $\text{O}^{2-}$  because it has the highest electron affinity  
 (5) all of the above are the same size because they are all isoelectronic
- What is the root mean square speed of  $\text{N}_2$  gas at STP?  
 (1) 15.6 cm/s (2) 243 cm/s (3)  $1.24 \times 10^4 \text{ cm/s}$  (4)  $4.93 \times 10^2 \text{ cm/s}$  (5) 893 cm/s
- A liquid has a normal boiling point of  $78^\circ\text{C}$  and a vapor pressure of 400 Torr at  $50^\circ\text{C}$ . What is the heat of vaporization of this liquid?  
 (1) 2.0 kJ/mol (2) 9.4 kJ/mol (3) 22 kJ/mol (4) 34 kJ/mol (5) 210 kJ/mol
- At  $20^\circ\text{C}$ , the vapor pressure of pure benzene is 75 torr and pure toluene is 22 torr. An equimolar solution of benzene and toluene is prepared. What is the mole fraction of benzene in the vapor above this solution at  $20^\circ\text{C}$ ?  
 (1) 0.23 (2) 0.29 (3) 0.50 (4) 0.77 (5) 0.83
- A sample of liquid water is electrolyzed, completely converting the water to hydrogen gas and oxygen gas which is collected in a closed container. If the total pressure in the container is 600 torr, what is the partial pressure of the hydrogen in the container?  
 (1) 100 torr (2) 200 torr (3) 300 torr (4) 400 torr (5) 500 torr
- The molecular shape of the  $\text{IBr}_3$  molecule is  
 (1) trigonal bipyramidal (2) seesaw (3) tee-shaped (4) trigonal pyramidal (5) trigonal planar
- Which of the following processes or reactions is not exothermic?  
 (1)  $\text{K}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{KBr}(\text{s})$  (2)  $\text{CO}_2(\text{g}) \rightarrow \text{CO}_2(\text{s})$  (3)  $\text{F}_2(\text{g}) \rightarrow 2\text{F}(\text{g})$   
 (4)  $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$  (5)  $\text{F}(\text{g}) + e^- \rightarrow \text{F}^-(\text{g})$
- Which of the following groups is arranged Left to Right in order of increasing atomic size?  
 (1) Mg, Na, K (2) K, Ca, Sr (3) Na, Mg, Al (4) Rb, K, Na (5) Br, Cl, F
- A first-order reaction has a rate constant of  $1.32 \times 10^{-3} \text{ yr}^{-1}$ . What is the half-life of the reaction?  
 (1)  $1.32 \times 10^{-3} \text{ yr}$  (2)  $1.90 \times 10^{-3} \text{ yr}$  (3) 525 yr (4) 758 yr (5) None of these

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**G Orfield**



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