

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

### Midterm Exam Study Guide

d. Isotopes

- i. Isotopes differ in their \_\_\_\_\_ and number of \_\_\_\_\_

e. Average atomic mass Formula:

f. Ions

- i. \_\_\_\_\_ have lost electrons and have a \_\_\_\_\_ charge  
ii. \_\_\_\_\_ have gained electrons and have a \_\_\_\_\_ charge

g. Group Numbers be able to locate them

- i. Alkali Metals  
ii. Alkaline earth Metals  
iii. Transition Metals  
iv. Inner Transition Metals  
v. Halogens  
vi. Noble Gases

h. Periodic Trends Give your notes for reference (be able to for each of these)

- i. Atomic Radius/Size  
ii. Ionic Radius/Size  
iii. Ionization Energy  
iv. Electronegativity

i. Valence Electrons be able to identify # of valence electrons and charge for each group

- i. Lewis-Dot structures

2. Nuclear Decay and Half-life

- a.  ${}^4_2\text{He}$  or  ${}^4_2\text{He}$  is called an \_\_\_\_\_ particle which has the weakest radiation.  
b.  ${}^0_{-1}\text{e}$  or  ${}^0_{-1}\beta$  is called an \_\_\_\_\_ particle which has moderate radiation and is an electron.  
c.  ${}^1_0\text{n}$  is called an \_\_\_\_\_ particle which has the strongest radiation and highest energy.  
d. Half-life  
i. Definition: \_\_\_\_\_  
ii. Know how to read a graph and calculate half-life

3. Electron Configurations

- a. Be able to draw them/identify them from another student  
b. Hund's Rule  
c. Pauli Exclusion Principle  
d. Aufbau Principle

4. Ionic Compounds

- a. Ionic compounds are made of \_\_\_\_\_ + \_\_\_\_\_  
b. Their overall charge is \_\_\_\_\_  
c. Know how to use Lewis-Dot Structures to draw bonds between cation and anion

# Mcps Chemistry Semester Review Guide Answers

**LM Reder**



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**International** ,2004      *Chemistry - the Physical Setting* ,2004      **AP Chemistry Exam Success** Lewis Morris,2018-07-14

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*Study Guide for Chemistry* H LeMay, Theodore Brown, Bruce Bursten, Catherine Murphy, Patrick Woodward, Matthew Stoltzfus, 2017-03-21 This guide assists students through the text material with chapter overviews learning objectives a review of key terms as well as self tests with answers and explanations This student guide also features MCAT practice questions

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*AP Chemistry Test Secrets*, 2010

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