

Page 3	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2013	0620	32

- 3 (a) (i) the (forward) reaction is endothermic [1]
- (ii) none [1]
volume of reactants and products the same [1]
ACCEPT: number of moles or molecules
- (iii) the reaction (between oxygen and nitric oxide) is **exothermic** [1]
high temperatures push equilibrium to left / high temperatures decrease yield of products [1]
/ low temperatures favour forward reaction
- (iv) $4\text{HNO}_3 + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + 4\text{HNO}_2$ [2]
not balanced = (1) only
- (v) (cost of) high amount of electricity / energy [1]
- (b) (i) contains more nitrogen [1]
- (ii) photosynthesis [1]
chlorophyll is catalyst / chlorophyll absorbs light [1]
carbon dioxide and water react [1]
to make glucose / carbohydrates / starch / sugar / named sugar [1]

[Total: 13]

- 4 (a) Any one of:
 $\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 2\text{Fe} + 3\text{CO}$
 $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
 $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
 for correct equation (2)
 not balanced = (1) only
- any four of:
- coke burns to form carbon dioxide / $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- this reacts with more carbon to form carbon monoxide / $\text{C} + \text{CO}_2 \rightarrow 2\text{CO}$
- calcium carbonate decomposes to form calcium oxide and carbon dioxide / $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
- calcium oxide / calcium carbonate reacts with silica / silicon oxide / silicon(IV) oxide (in one)
 to form calcium silicate / slag / $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$ or $\text{CaCO}_3 + \text{SiO}_2 \rightarrow \text{CaSiO}_3 + \text{CO}_2$
- the reaction between carbon and oxygen is exothermic / produces heat / coke is used as a fuel / the slag floats on the (molten) iron / the slag and molten iron can be run off separately [2]

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