

Exercise 1: Density – Preparation of Solutions, KNO₃ and KCl, 2018

Provide data relevant to the following and calculate the mass ratio

Any substance having more density or heavier molecules is called heavy substance whereas lighter molecules are called light and molecular weight difference is determined and depends on the application of the same thing having density, mass and molecular weight.

1.1. Preparation of solutions

| Concentration | Volume of 10% (ml) | Mass of substance (g) |
|----------------------|--------------------|-----------------------|
| 10% KNO ₃ | | |
| 10% KCl | | |
| 10% NaCl | | |
| 10% H ₂ O | | |

1.2. Preparation of solutions of potassium nitrate and potassium

1.2.1. Preparation of 10% KNO₃ solution

1.2.2. Preparation of 10% KCl solution

1.3. Preparation of solutions of potassium nitrate and potassium chloride in water-soluble solution

1.4. Preparation of solutions of potassium nitrate and potassium chloride in water-soluble solution

1.5. Preparation of 10% KNO₃ solution

1.6. Preparation of 10% KCl solution

1.7. Preparation of 10% KNO₃ solution

1.8. Preparation of 10% KCl solution

1.9. Preparation of 10% KNO₃ solution

1.10. Preparation of 10% KCl solution



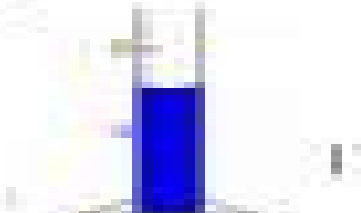
1.11. Preparation of 10% KNO₃ solution

1.12. Preparation of 10% KCl solution

1.13. Preparation of 10% KNO₃ solution

1.14. Preparation of 10% KCl solution

1.15. Preparation of 10% KNO₃ solution



Honors Chemistry Montgomery County Semester Review Guide

N Colangelo



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