

1. Determine if a pure substance is made of atoms or molecules.

_____ C	_____ H ₂ O	_____ carbon dioxide	_____ sodium
_____ Ar	_____ Cu	_____ bromine	_____ sulfur
_____ PBr ₃	_____ P ₄	_____ sulfur trioxide	_____ neon
_____ Ca	_____ N ₂	_____ silicon	_____ fluorine
_____ H ₂	_____ C ₆ H ₁₂ O ₆	_____ boron trioxide	_____ iron

2. Distinguish between pure substances and mixtures.

_____ NaCl & H ₂ O	_____ sand & sugar
_____ CO ₂ & O ₂	_____ gold
_____ CH ₄	_____ lead (II) nitrate dissolved in water
_____ 2n(C ₆ H ₅ O ₂) ₂ (aq)	_____ nitrogen trihydride and oxygen
_____ Ca ²⁺ & SO ₄ ²⁻ & H ₂ O	_____ marbles and beads

3. Distinguish between elements and compounds.

_____ I ₂	_____ nitrogen monoxide
_____ Ag ₃ P	_____ nickel
_____ Sr ₃ N ₂	_____ chromium (III) sulfide
_____ CO	_____ aluminum bromide
_____ Cu	_____ fluorine

4. Distinguish between homogeneous mixtures (solutions) and heterogeneous mixtures.

- i. How can you tell if a mixture is homogeneous or heterogeneous?

5. Describe a chemical on the macroscopic, particle, and symbolic levels.

- i. Know how to represent a chemical with particle pictures.
ii. Know how to represent a chemical with symbols (formulas).

6. Know which elements are diatomic in nature and which are polyatomic in nature.

- i. Know the name and symbol of each diatomic (7) and polyatomic element (2) in nature.

7. Know the names and symbols of the first 54 elements on the periodic table.

- i. Be familiar with the first 54 elements so you can easily locate it on the periodic table.
ii. If you have them memorized, you will save time having to look it up on the periodic table.

8. Understand how to use the periodic table to identify groups/families of elements (and their chemical properties) known as alkali metals, alkaline earth metals, halogens, and noble gases.

- i. Which groups are alkali metals, alkaline earth metals, halogens, and noble gases found?

9. Distinguish between metals and nonmetals.

_____ H	_____ Cl	_____ lithium	_____ silver
_____ N	_____ Cr	_____ oxygen	_____ carbon
_____ Sr	_____ P	_____ arsenic	_____ sulfur

10. Understand the nature of the covalent bond and ionic bond.

- i. Covalent bonds are formed when two nonmetals fight over or share electrons.
ii. Ionic bonds are formed when a cation (being positively charged because it lost electrons) is attracted to an anion (being negatively charged because it gained electrons).
a. Ionic formulas are written in simplest whole number ratios.

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