

HEAT AND ITS MEASUREMENT

Name _____

Heat (or energy) can be measured in units of calories or joules. When there is a temperature change (ΔT), heat (Q) can be calculated using this formula:

$$Q = \text{mass} \times \Delta T \times \text{specific heat capacity}$$

$$(\Delta T = \text{Final Temp} - \text{Initial Temp})$$

During a phase change, we use this formula:

$$Q = \text{mass} \times \text{heat of fusion (or heat of vaporization)}$$

Solve the following problems.

1. How many joules of heat are given off when 5.0 g of water cool from 75° C to 25° C? (Specific heat of water = 4.18 J/g° C)

$$(5)(50)(4.18) = 1045 \text{ rounded } 1000 \text{ J}$$

2. How many calories are given off by the water in Problem 1? (Specific heat of water = 1.0 cal/g° C)

$$(5)(50)(1) = 250 \text{ cal}$$

3. How many joules does it take to melt 35 g of ice at 0° C? (heat of fusion = 333 J/g)

$$(35g)(333) = 11655$$

No ΔT when melting Just 0° 12000 J

4. How many calories are given off when 85 g of steam condense to liquid water? (heat of vaporization = 539.4 cal/g)

$$46000 \text{ cal}$$

5. How many joules of heat are necessary to raise the temperature of 25 g of water from 10° C to 60° C?

$$5000 \text{ J}$$

6. How many calories are given off when 50 g of water at 0° freezes? (heat of fusion = 79.72 cal/g)

$$4000 \text{ cal}$$

Instructional Fair Worksheets Heat And Its Measurement Answers

PT Brinkman



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Table of Contents Instructional Fair Worksheets Heat And Its Measurement Answers

1. Understanding the eBook Instructional Fair Worksheets Heat And Its Measurement Answers
 - The Rise of Digital Reading Instructional Fair Worksheets Heat And Its Measurement Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Instructional Fair Worksheets Heat And Its Measurement Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Instructional Fair Worksheets Heat And Its Measurement Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Instructional Fair Worksheets Heat And Its Measurement Answers
 - Personalized Recommendations
 - Instructional Fair Worksheets Heat And Its Measurement Answers User Reviews and Ratings
 - Instructional Fair Worksheets Heat And Its Measurement Answers and Bestseller Lists

5. Accessing Instructional Fair Worksheets Heat And Its Measurement Answers Free and Paid eBooks
 - Instructional Fair Worksheets Heat And Its Measurement Answers Public Domain eBooks
 - Instructional Fair Worksheets Heat And Its Measurement Answers eBook Subscription Services
 - Instructional Fair Worksheets Heat And Its Measurement Answers Budget-Friendly Options
6. Navigating Instructional Fair Worksheets Heat And Its Measurement Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Instructional Fair Worksheets Heat And Its Measurement Answers Compatibility with Devices
 - Instructional Fair Worksheets Heat And Its Measurement Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Instructional Fair Worksheets Heat And Its Measurement Answers
 - Highlighting and Note-Taking Instructional Fair Worksheets Heat And Its Measurement Answers
 - Interactive Elements Instructional Fair Worksheets Heat And Its Measurement Answers
8. Staying Engaged with Instructional Fair Worksheets Heat And Its Measurement Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Instructional Fair Worksheets Heat And Its Measurement Answers
9. Balancing eBooks and Physical Books Instructional Fair Worksheets Heat And Its Measurement Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Instructional Fair Worksheets Heat And Its Measurement Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Instructional Fair Worksheets Heat And Its Measurement Answers
 - Setting Reading Goals Instructional Fair Worksheets Heat And Its Measurement Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Instructional Fair Worksheets Heat And Its Measurement Answers
 - Fact-Checking eBook Content of Instructional Fair Worksheets Heat And Its Measurement Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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