

ATOMIC STRUCTURE

Name _____

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.

Example:

mass
number
↓
15
↑
atomic
number
7

N

charge
↖
+3

7 protons
8 neutrons (15 - 7)
4 electrons

Always write the magnitude of the charge followed by the sign: N3+

Can you find the other mistakes on this page?

Complete the following chart.

Element/ Ion	Atomic Number	Atomic Mass	Mass Number	Protons	Neutrons	Electrons
H						
H ⁺						
¹² ₆ C						
⁷ ₃ Li ⁺						
³⁵ ₁₇ Cl ⁻						
³⁹ ₁₉ K						
²⁴ ₁₂ Mg ²⁺						
As ³⁻						
Ag						
Ag ⁺¹						
S ⁻²						
U						

Instructional Fair Inc Parts Of An Atom

D Siedentop



Instructional Fair Inc Parts Of An Atom:

Decoding **Instructional Fair Inc Parts Of An Atom**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Instructional Fair Inc Parts Of An Atom**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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