

GRADE 12 PRESCRIBED EXPERIMENT 2: ACID-BASE REACTIONS
WORK SHEET
TOTAL MARK: 50

ACTIVITY

Titration of oxalic acid against sodium hydroxide to determine the concentration of the sodium hydroxide.

In this investigation you will prepare an acidic solution accurately and thus you will know its exact concentration. You will then react this acid with a base of an unknown concentration to determine the concentration of the base.

What you will need:

Erlenmeyer flasks	Oxalic acid
Burettes	Sodium Hydroxide
Burette clamp	Phenolphthalein as indicator
Medicine dropper	Funnel
Retort stand	Beaker
White tile /paper	Spatula
Measuring cylinders	Glass rod
Mass meter	Pipette with sucker

What to do:

1. Prepare a standard solution of oxalic acid which has a concentration of approximately 1 mol dm^{-3} .
2. Now prepare a sodium hydroxide solution by dissolving approximately 2 g of dry sodium hydroxide in 500 ml of water.
3. Add two drops of the indicator solution.
4. Place the burette in the clamp.
5. Using the funnel, fill the burette to above zero mark with the acid solution.
6. Then, holding the beaker, with which you used to pour the acid, beneath the burette, gradually open the tap.
7. Allow the level of the base to come down to exactly zero (reading from the bottom of the meniscus).
8. Pipette using the sucker exactly 25 ml of oxalic acid solution in a volumetric flask.
9. Add a few drops of phenolphthalein to the acid.
10. Hold the conical flask beneath the burette with your right hand and gradually open the tap with your left.
11. Swirl the conical flask continuously and watch it closely for the first sign of a colour change.
12. As you see that you are approaching the point of neutralization, close the tap slightly so that you are adding drop by drop.
13. When the colour changes completely the titration is finished.
14. Close the tap and read from the burette how much acid was used.
15. Repeat this procedure at least twice so that you have three readings for the volume

Grade 12 Prescribed Experiment 2

AN Whitehead



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