

CHEMISTRY

22/06/2022

CONTENT:

- Quantitative chemistry
- Chemical bonding
- Energy Change & Rate of reactions
- Chemical Equilibrium
- Organic chemistry

QUANTITATIVE CHEMISTRY:

molar mass: mass in grams of one mole of that substance.

$$n = \frac{m}{M}$$

number of mol $\leftarrow$ mass (g) $\leftarrow$ Molar mass

STP: Standard temperature and pressure

At STP, 1 mole of any gas occupies 22.4 dm³

$$n = \frac{V}{V_m}$$

number of mol $\leftarrow$ Volume $\leftarrow$ 22.4 (constant)

NOTE:
1 cm³ = 1 ml
1 dm³ = 1 L

Concentration of solutions:

solution: homogeneous mixture of solute and solvent

Solute: Substance that is dissolved in solution

Solvent: Substance in which another substance is dissolved, forming a solution.

concentration: Amount of solute per unit volume of solution.

$$C = \frac{n}{V}$$

concentration $\leftarrow$ number of mol solute $\leftarrow$ volume of solution

$$C = \frac{m}{M \cdot V}$$

concentration $\leftarrow$ mass (g) $\leftarrow$ molar mass $\leftarrow$ volume

Stoichiometric calculations:

Balancing numbers give mole ratio.

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