

**QUESTION 2**

- 2.1 The graph of molecular size versus the boiling point is given below.  
The letters **A**, **B** and **C** represent the compounds  $\text{CH}_4$ ,  $\text{C}_2\text{H}_6$  and  $\text{C}_3\text{H}_8$  respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ( $^{\circ}\text{C}$ )
Ammonia	$\text{NH}_3$	17	- 33
Phosphine	$\text{PH}_3$	x	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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  - The Rise of Digital Reading Grade11 June Exam Paper1 2012 Physical Science
  - Advantages of eBooks Over Traditional Books
2. Identifying Grade11 June Exam Paper1 2012 Physical Science
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade11 June Exam Paper1 2012 Physical Science
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grade11 June Exam Paper1 2012 Physical Science

- Personalized Recommendations
  - Grade11 June Exam Paper1 2012 Physical Science User Reviews and Ratings
  - Grade11 June Exam Paper1 2012 Physical Science and Bestseller Lists
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- Grade11 June Exam Paper1 2012 Physical Science Public Domain eBooks
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  - Grade11 June Exam Paper1 2012 Physical Science Budget-Friendly Options
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- ePub, PDF, MOBI, and More
  - Grade11 June Exam Paper1 2012 Physical Science Compatibility with Devices
  - Grade11 June Exam Paper1 2012 Physical Science Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Grade11 June Exam Paper1 2012 Physical Science
  - Highlighting and Note-Taking Grade11 June Exam Paper1 2012 Physical Science
  - Interactive Elements Grade11 June Exam Paper1 2012 Physical Science
8. Staying Engaged with Grade11 June Exam Paper1 2012 Physical Science
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade11 June Exam Paper1 2012 Physical Science
9. Balancing eBooks and Physical Books Grade11 June Exam Paper1 2012 Physical Science
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grade11 June Exam Paper1 2012 Physical Science
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Grade11 June Exam Paper1 2012 Physical Science
- Setting Reading Goals Grade11 June Exam Paper1 2012 Physical Science
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade11 June Exam Paper1 2012 Physical Science

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  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
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  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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