

10.1**THE NATURE OF GASES****SECTION REVIEW****Objectives**

- Describe the motion of gas particles according to the kinetic theory
- Interpret gas pressures in terms of kinetic theory

Key Terms

- kinetic energy
- kinetic theory
- pressure (Pa)
- gas pressure
- vacuum
- standard atmospheric pressure
- atmospheric pressure
- barometric pressure

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are highlighted in this section. Write the blank over the completed word or phrase, about phrases, or names here.

- The kinetic theory describes the _____ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly _____, that the particles of gas are _____, move _____ of each other, and are in constant _____ motion. Furthermore, the _____ between particles are elastic so that the total _____ remains constant. The pressure of a gas results from the countless collisions of the gas particles with an object. Barometers are used to measure _____ pressure. Standard conditions are defined as a temperature of _____ and a pressure of _____.
- _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____
 - _____

Part B True-False

Classify each of these statements as true (always true), AT: sometimes true, ST: answer true, NT: answer false.

- _____ 11. Atmospheric pressure is 760 mm Hg.
- _____ 12. The SI unit of pressure is the pascal.
- _____ 13. Atmospheric pressure decreases as you climb a mountain because the air layer that surrounds the earth thins out.

Nature Of Gases Section Review Answer Key

E Durkheim



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