

# MOLARITY BY DILUTION

79 + 52 = 131 Name \_\_\_\_\_

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid =  $M \times V$  then,  
 $M_1 \times V_1 = M_2 \times V_2$  Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(6.0 M)(250 \text{ mL})}{18 M}$$

83 mL

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100 mL of a 2.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(2.0 M)(100 \text{ mL})}{12 M}$$

17 mL

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(15 M)(25 \text{ mL})}{3.0 M}$$

125 mL

4. To how much water should 50. mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(12 M)(50 \text{ mL})}{4.0 M} = 150 \text{ mL} - 50 \text{ mL}$$

100 mL H<sub>2</sub>O

5. To how much water should 100. mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(18 M)(100 \text{ mL})}{1.5 M} = 1200 \text{ mL} - 100 \text{ mL} = 1100 \text{ mL}$$

1.1 L

# Instructional Fair In Molarity Worksheet Answer Key

**Frank Schaffer Publications, Joan  
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## **Table of Contents Instructional Fair In Molarity Worksheet Answer Key**

1. Understanding the eBook Instructional Fair In Molarity Worksheet Answer Key
  - The Rise of Digital Reading Instructional Fair In Molarity Worksheet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Instructional Fair In Molarity Worksheet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Instructional Fair In Molarity Worksheet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Instructional Fair In Molarity Worksheet Answer Key
  - Personalized Recommendations
  - Instructional Fair In Molarity Worksheet Answer Key User Reviews and Ratings
  - Instructional Fair In Molarity Worksheet Answer Key and Bestseller Lists
5. Accessing Instructional Fair In Molarity Worksheet Answer Key Free and Paid eBooks
  - Instructional Fair In Molarity Worksheet Answer Key Public Domain eBooks
  - Instructional Fair In Molarity Worksheet Answer Key eBook Subscription Services
  - Instructional Fair In Molarity Worksheet Answer Key Budget-Friendly Options
6. Navigating Instructional Fair In Molarity Worksheet Answer Key eBook Formats

- ePub, PDF, MOBI, and More
  - Instructional Fair In Molarity Worksheet Answer Key Compatibility with Devices
  - Instructional Fair In Molarity Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Instructional Fair In Molarity Worksheet Answer Key
    - Highlighting and Note-Taking Instructional Fair In Molarity Worksheet Answer Key
    - Interactive Elements Instructional Fair In Molarity Worksheet Answer Key
  8. Staying Engaged with Instructional Fair In Molarity Worksheet Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Instructional Fair In Molarity Worksheet Answer Key
  9. Balancing eBooks and Physical Books Instructional Fair In Molarity Worksheet Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Instructional Fair In Molarity Worksheet Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Instructional Fair In Molarity Worksheet Answer Key
    - Setting Reading Goals Instructional Fair In Molarity Worksheet Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Instructional Fair In Molarity Worksheet Answer Key
    - Fact-Checking eBook Content of Instructional Fair In Molarity Worksheet Answer Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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